

Chapter 7 Chemical Reactions

Section 7.5 Equilibrium**(pages 216–219)**

This section explains physical and chemical equilibria, and describes the factors that affect chemical equilibrium.

Reading Strategy (page 216)

Outlining As you read, make an outline of the most important ideas from this section. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook.

I. Equilibrium**A. Types of Equilibria**

1.

2.

B.

1. Temperature

2. Pressure

3.

Types of Equilibria (pages 216–217)

1. What is equilibrium? _____

2. Circle the letter of the correct answer. In the system described by the equation $\text{H}_2\text{O}(l) \rightleftharpoons \text{H}_2\text{O}(g)$, at room temperature, which of the following two physical changes are in equilibrium?

- a. sublimation and condensation
- b. evaporation and melting
- c. sublimation and deposition
- d. evaporation and condensation

3. What happens when a physical change does not go to completion? _____

4. What does the single arrow imply about the reaction described in the following equation?



7 5 Equilibrium Answers

Gary S Thorpe



7 5 Equilibrium Answers:

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Power Practice: Physical Science, eBook Beth Barber, 2004-09-01 This book supplements and enriches classroom teaching to enhance students' understanding of vocabulary functions and fundamental processes of physical sciences work Topics include force and motion chemistry atoms and elements scientific process simple machines energy light and sound magnetism and electricity

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