1989 Advanced Placement Exam

Part I Multiple Choice

Note: For all questions referring to solutions, assume that the solvent is water unless otherwise stated.

<u>Directions</u>: Each set of lettered choices below refers to the numbered statements immediately following it. Select the one lettered choice that best fits each statement and then blacken the corresponding space on the answer sheet. A choice may be used once, more than once, or not at all in each set.

Ouestions 1-3

a. O	c. Rb e	n N
b. La	d. Mg	

- What is the most electronegative element?
- Which element exhibits the greatest number of different oxidation states?
- 3. Which of the elements above has the smallest ionic radius for its most commonly found ion?

Ouestions 4-7

\mathbf{a} . $1 \mathbf{s}^{\mathbf{r}}$	$2s^{3}2p^{6} - 3s^{3}3p^{6}$
b. $1s^{\circ}$	$2s^{2}2p^{\mu} - 3s^{2}3p^{\mu}$
c. 1.s ²	$2s^{2}2p^{p}2d^{p}-3s^{2}3p^{p}$
$d. 1s^{s}$	2s*2p* 3s*3p*3d*
	25 20 35 30 30 45

- An impossible electronic configuration
- The ground-state configuration for the atoms of a 5 transition element
- The ground-state configuration of a negative ion Questions 11-14 of a halogen
- The ground-state configuration of a common ion 7... of an alkaline earth element

Questions 8-10

- a. a solution with a pH less than 7 that is not a buffer solution
- a buffer solution with a pH between 4 and 7
- c. a buffer solution with a pH between 7 and
- d. a solution with a pH greater than 7 that is not a buffer solution
- e. a solution with a pH of 7

(Ionization constants: CH₃COOH = 1.8∞10⁻⁶; $NH_b = 1.8 \times 10^{-5}$. H_2CO_3 : $K_1 = -4 \times 10^{-4}$; $K_{*} = -4 \times 10^{-0}$;)

- 8. A solution prepared to be initially 1 M in NaCl and I M in HCL
- 9. A solution prepared to be initially 1 M in Na₂CO₃ and 1 M in CH₂COONa
- 10. A solution prepared to be initially 0.5 M in CH-COOH and 1 M in CH-COONa

- a. hydrogen bonding
- b. hybridization
- e. ionie bonding
- d. resonance
- e. van der Waals forces (London dispersion forces).
- 11. Is used to explain why iodine molecules are held together in the solid state
- 12. Is used to explain why the boiling point of HF is greater than the boiling point of HBr
- Is used to explain the fact that the four bonds in methane are equivalent
- 14. Is used to explain the fact that the carbon-tocarbon bonds in benzene, C.H., are identical

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