
(General Chemistry Answer Key)

1. What is the atomic number of carbon?

Answer: 6

2. What is the molar mass of H_2O ?

Answer: 18.02 g/mol

3. How many moles are in 36.0 grams of water (H_2O)?

Answer: 2 moles (36.0 g \div 18.02 g/mol)

4. What is the empirical formula of $\text{C}_8\text{H}_{10}\text{O}_2$?

Answer: CH_2O

5. Balance the following chemical equation:



Answer: $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$

6. What type of bond is formed between sodium (Na) and chlorine (Cl)?

Answer: ionic bond

7. What is the molarity of a solution prepared by dissolving 58.5 grams of NaCl (MW = 58.5 g/mol) in 1 liter of water?

Answer: 1 M (58.5 g \div 58.5 g/mol = 1 mol in 1 L)

8. What is the pH of a solution with $[\text{H}^+] = 1 \times 10^{-7} \text{ M}$?

Answer: pH = 7 (since $\text{pH} = -\log[\text{H}^+]$)

9. What is the chemical formula of calcium phosphate?

Answer: $\text{Ca}_3(\text{PO}_4)_2$

10. How many atoms are in 2.5 moles of nitrogen gas (N_2)?

Answer: 3.0×10^{24} atoms (2.5 moles \times 6.022×10^{23} atoms/mol)

11. What is the oxidation state of sulfur in H_2SO_4 ?

Answer: +6

12. What is the limiting reactant in the following reaction:



Answer: O_2 (since 2 moles of H_2 require 1 mole of O_2 , and O_2 runs out first)

13. What is the enthalpy change (ΔH) for the reaction:



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Peter May



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