12-3 Practice Problems

1. From the following enthalpy changes,

 $C(x) + \frac{1}{2}O_2(g) \rightarrow CO(g)$

 $\Delta H^* = -110.5 \text{ kJ}$

 $CO(g) + \frac{1}{2}O_2(g) \rightarrow CO_2(g)$

 $\Delta H^+ = -283.0 \text{ kJ}$

calculate the value of ΔH^* for the reaction $C(s) + O_2(g) \rightarrow CO_2(g)$.

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2. From the following enthalpy changes,

 $2\mathbb{P}(s) + 3\mathbb{Cl}_2(g) \to 2\mathbb{P}\mathbb{Cl}_3(t)$

 $\Delta H^* = -640 \text{ kJ}$

 $2P(s) + 5Cl_s(g) \rightarrow 2PCl_c(s)$

 $\Delta H^{+} = -886 \text{ kJ}$

calculate the value of ΔH^+ for the reaction $[PCL_t(t) + CL_t(g)] \rightarrow PCL_t(s)$

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Prom the following enthalpy changes,

 $C_2H_5OH(l) + 3O_2(g) \rightarrow 2CO_2(g) + 3H_2O(g)$ $\Delta H^+ = -1234.7 \text{ kJ}$

 $CH_3OCH_3(I) + 3O_2(g) \rightarrow 2CO_2(g) + 3H_2O(g)$ $\Delta H^a = -1328.3 \text{ kJ}$

calculate the value of ΔH° for the reaction $C_2H_3OH(l) \rightarrow CH_3OCH_3(l)$.

From the following enthalpy changes,

 $Cu(s) + Cl_2(g) \rightarrow CuCl_2(s)$

 $\Delta H^{\circ} = -206 \text{ kJ}$

 $2Cu(s) + Cl_2(g) \rightarrow 2CuCl(s)$

 $\Delta H^{\circ} = -136 \text{ kJ}$

calculate the value of ΔH^* for the reaction $CuCl_2(s) + Cu(s) \rightarrow 2CuCl(s)$. L = 0 L =

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5. From the following enthalpy changes,

 $H_2(g) + F_2(g) \rightarrow 2HF(g)$ $\Delta H^0 = -542.2 \text{ kJ}$

 $2H_2(g) + O_2(g) \rightarrow 2H_2O(l)$ $\Delta H^* = -571.6 \text{ kJ}$

calculate the value of ΔH^* for the reaction $2F_2(g) + 2H_2O(I) \rightarrow 4HF(g) + O_2(g)$

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6. From the following enthalpy changes,

 $Xe(g) + F_g(g) \rightarrow XeF_g(s)$

 $\Delta H^* = -123 \text{ kJ}$

 $Xe(g) + 2F_2(g) \rightarrow XeF_4(s)$

 $\Delta H^* = -262 \text{ kJ}$

calculate the value of ΔH^a for the reaction $XeF_a(s) + F_a(g) \rightarrow XeF_a(s)$.

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7. From the following enthalpy changes,

 $4NH_3(g) + 5O_2(g) \rightarrow 4NO(g) + 6H_2O(I)$ $\Delta H^* = -1170 \text{ kJ}$

 $4NH_3(g) + 3O_2(g) \rightarrow 2N_2(g) + 6H_2O(l)$ $\Delta H^* = -1530 \text{ kJ}$

calculate the value of ΔH^* for the reaction $N_2(g) + O_2(g) \rightarrow 2NO(g)$

8. From the following enthalpy changes,

 $2Al(s) + \frac{3}{2} O_2(g) \rightarrow Al_2O_3(s)$ $\Delta H^0 = -1601 \text{ kJ}$ $2Fe(s) + \frac{3}{2} O_2(g) \rightarrow Fe_2O_3(s)$ $\Delta H^0 = -821 \text{ kJ}$

calculate the value of ΔH^{+} for the reaction $2Al(s) + Fe_{2}O_{3}(s) \rightarrow 2Fe(s) + Al_{3}O_{3}(s)$

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9. From the following enthalpy changes,

 $H_2O_2(I) \rightarrow H_2O(I) + \frac{1}{2}O_2(g)$

 $\Delta H^* = -94.6 \text{ kJ}$ $\Delta H^* = -286.0 \text{ kJ}$

 $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(l)$ $\Delta H^* = -286.0 \text{ kJ}$ calculate the value of ΔH^* for the reaction

 $H_2(g) + H_2O_2(I) \rightarrow 2H_2O(I)$ $H_2(g) + H_2O_2(I) \rightarrow 2H_2O(I)$ $H_1 + H_2O_2 + H$

10. From the following enthalpy changes,

 $C(s) + O_2(g) \rightarrow CO_2(g)$

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 $\Delta H^{\circ} = -393.5 \text{ kJ}$

 $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(f)$ $\Delta H^* = -285.8 \text{ kJ}$ $2C_2H_2(g) + 5O_2(g) \rightarrow 4CO_2(g) + 2H_2O(f)$

ΔH* = −2598.5 kJ

calculate the value of ΔH^* for the reaction 2C(s)+ $H_2(g) \rightarrow C_2H_2(g)$. $2\zeta \cdot R_1 \rightarrow \zeta A_1$

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