

1. Predict the order of increasing vapor pressure at a given temperature for the following compounds. 1 = lowest vapor pressure at a given temperature. (i.e. the liquid with the lowest boiling point will have more molecules in the gas phase at any temperature below its boiling point and thus will have a higher vapor pressure. The pressure exerted by the molecules in the gas phase is the vapor pressure). Explain your choice for highest and lowest vapor pressure.

lowest boiling pt = highest vapor pressure

$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$ $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$ $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$ $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

3 1 2 in gas phase at any temp since it forms weakest intermolecular attractions (dipole-dipole).

$\text{H}_2\text{O}-\text{CH}_2-\text{CH}_2-\text{OH}$ will be lowest VP since it forms H-bonds at both ends of molecule

2. Which of the following will occur if the temperature of a gas is increased from 15 °C to 55 °C at constant pressure?

- (a) the volume remain the same (b) the volume will increase (c) the volume will decrease

3. A gas occupies 2.7 L at 0.40 atm pressure. If we expand the volume at constant temperature to 4.1 L, what is the final pressure? CIRCLE your final answer.

$$P_1 V_1 = P_2 V_2 \quad P_2 = \frac{P_1 V_1}{V_2} = \frac{(0.40 \text{ atm})(2.7 \text{ L})}{4.1 \text{ L}} = 0.26 \text{ atm}$$

4. A vessel under 3.08 atm pressure contains carbon dioxide and water vapor. The partial pressure of water is 2.21 atm. What is the partial pressure of carbon dioxide? CIRCLE your final answer.

$$P_T = 3.08 \text{ atm} = P_{\text{CO}_2} + P_{\text{H}_2\text{O}} \quad P_{\text{CO}_2} = 3.08 \text{ atm} - 2.21 \text{ atm} = 0.87 \text{ atm}$$

5. Dry air is 78.08% nitrogen, 20.95% oxygen, and 0.93% argon with the 0.04% being other gases. If the atmospheric pressure is 760.0 torr, what is the partial pressure of argon in dry air? CIRCLE your final answer.

$$(760.0 \text{ torr})(0.0093) = 7.07 \text{ torr}$$

6. A constant volume of oxygen gas is heated from 115 °C to 225 °C. The final pressure is 17.7 atm. What was the initial pressure? CIRCLE your final answer.

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad P_1 = \frac{P_2 T_1}{T_2} = \frac{(17.7 \text{ atm})(388 \text{ K})}{498 \text{ K}} = 13.8 \text{ atm}$$

7. A balloon has an initial volume of 2.25 L. What will the volume be if the temperature is raised from 20 °C to 35 °C?

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad V_2 = \frac{V_1 T_2}{T_1} = \frac{(2.25 \text{ L})(308 \text{ K})}{293 \text{ K}} = 2.37 \text{ L}$$

8. A bicycle tire is inflated to a total pressure of 125 psi. What is the pressure in mmHg?

- (a) 165 (b) 8,500 (c) 8.50 (d) 6.46 (e) 6.46×10^2

$$125 \text{ psi} \times \frac{760 \text{ mmHg}}{14.7 \text{ psi}} = 6462.5 \text{ mmHg}$$

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sir George Christopher T. Bartley



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