

# Colligative Properties Worksheet

Reference: Normal 000°

1°C

0.1°C (0.1°C) (Normal 000°)

1°C

0.1°C (0.1°C)

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0.1°C (0.1°C) (0.1°C) (0.1°C) (0.1°C)

1. What does normality of particles that contribute to the normality of a solution?

Normal: 1°C (0.1°C) (0.1°C) (0.1°C) (0.1°C)

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Normal: 1°C (0.1°C) (0.1°C) (0.1°C) (0.1°C)

2. What is the normality of a solution of 1.00 g of NaCl in 100 mL of water?

0.100 g of NaCl in 100 mL of water is 0.100 g of NaCl in 100 mL of water. The normality of a solution is the number of moles of solute per liter of solution. The normality of a solution is the number of moles of solute per liter of solution.

3. If you prepared 1.00 g of NaCl in 100 mL of water, what would be the boiling point elevation and the freezing point?

0.100 g of NaCl in 100 mL of water is 0.100 g of NaCl in 100 mL of water. The normality of a solution is the number of moles of solute per liter of solution. The normality of a solution is the number of moles of solute per liter of solution.

4. If you prepared 1.00 g of NaCl in 100 mL of water, what would be the boiling point?

0.100 g of NaCl

5. If you prepared 1.00 g of NaCl in 100 mL of water, what would be the boiling point?

0.100 g of NaCl

6. If you prepared 1.00 g of NaCl in 100 mL of water, what would be the boiling point?

0.100 g of NaCl

# 18 3 Colligative Properties Of Solutions Answer Key

**DP Hallahan**



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