

14-1 Review and Reinforcement

Condensed States of Matter

Fill in the following chart comparing solids and liquids.

| Physical Property | Liquids | Solids |
|----------------------|---------|--------|
| 1. Compressibility | | |
| 2. Density | | |
| 3. Shape | | |
| 4. Rate of Diffusion | | |

Answer each of the following questions in the space provided.

5. Describe the behavior of the molecules in a liquid. Explain this behavior in terms of intermolecular forces.
- _____
- _____
6. How do intermolecular forces determine molecular arrangement in solids?
- _____
- _____

If the statement is true, write "true." If it is false, change the underlined word or words to make the statement true. Write your answer on the line provided.

- _____ 7. Intermolecular forces determine metallic properties such as the boiling point of a substance.
- _____ 8. Intermolecular forces are forces of attraction between atoms.
- _____ 9. Chemical bonds are the forces between molecules.
- _____ 10. Covalent bonds result from electrons being shared between atoms in a molecule.
- _____ 11. Intermolecular forces result from the electron interactions between neighboring molecules.
- _____ 12. When the electron cloud is not distributed symmetrically, a molecule has a polarity.
- _____ 13. Chemical bonds include ionic bonds, metallic bonds, and atomic bonds.
- _____ 14. Intermolecular forces include dispersion forces, dipole-dipole forces, and helium-bond forces.

14 Review And Reinforcement Answers

R Barnett



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