

21.1

NEUTRALIZATION REACTIONS

SECTION REVIEW

Objectives

- Explain how acid–base titrations is used to calculate the concentration of an acid or a base.
- Explain the concept of equivalence in neutralization reactions.

Key Terms

- neutralization reactions
- standard solution
- titration
- end point
- equivalent weight
- gram equivalent mass
- normality (N)
- equivalent point

Key Equations

- $\text{Acid} + \text{Base} \rightarrow \text{Salt} + \text{Water}$
- Grams equivalent mass = $\frac{\text{molar mass}}{\text{number of H}^+ \text{ or OH}^- \text{ ions}}$
- Normality (N) = $\frac{\text{equivalents}}{\text{L}}$
- $M_1 V_1 = N_1 V_1 = M_2 V_2 = N_2 V_2$
- $M_1 \times V_1 = M_2 \times V_2$

Part A Completion

Use their completion exercises to check your understanding of the concepts and terms that are emphasized in this section. Write clearly on the completed work a correct answer, or a new term.

In the reaction of an _____ acid with a base, hydrogen ions and _____ ions react to produce _____. This reaction, called _____, is usually carried out by _____. The _____ is a titration is the point at which the solution is neutral. At the _____ point of a titration, the number of equivalents of acid equals the number of equivalents of base.

An equivalent of acid is the _____ of the acid that provides _____ of hydrogen ions in solution. A solution that contains one equivalent of an acid or a base in one liter of solution is a one normal (1N) solution. Thus, a 2N solution of H_2SO_4 contains _____ of H^+ per liter of solution.

- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____

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