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**2018 SCORING GUIDELINES**

**Question 1**



Elemental sulfur can exist as molecules with the formula  $S_8$ . The  $S_8$  molecule is represented by the incomplete Lewis diagram above.

- (a) The diagram of  $S_8$  shows only bonding pairs of electrons. How many lone pairs of electrons does each S atom in the molecule have?

|     |                                           |
|-----|-------------------------------------------|
| Two | 1 point is earned for the correct answer. |
|-----|-------------------------------------------|

- (b) Based on your answer to part (a), determine the expected value of the S–S–S bond angles in the  $S_8$  molecule.

|                                                                                                                                              |                                                            |
|----------------------------------------------------------------------------------------------------------------------------------------------|------------------------------------------------------------|
| $109.5^\circ$<br>Acceptable range: $104^\circ \leq \text{angle} \leq 110^\circ$ .<br>(The experimentally determined angle is $107.8^\circ$ ) | 1 point is earned for an angle<br>in the acceptable range. |
|----------------------------------------------------------------------------------------------------------------------------------------------|------------------------------------------------------------|

- (c) Write the electron configuration for the S atom in its ground state.

|                                                       |                                                |
|-------------------------------------------------------|------------------------------------------------|
| $1s^2 2s^2 2p^6 3s^2 3p^4$ OR $[\text{Ne}] 3s^2 3p^4$ | 1 point is earned for a correct configuration. |
|-------------------------------------------------------|------------------------------------------------|

- (d) The complete photoelectron spectrum for the element chlorine is represented below. Peak X in the spectrum corresponds to the binding energy of electrons in a certain orbital of chlorine atoms. The electrons in this orbital of chlorine have a binding energy of 273 MJ/mol, while the electrons in the same orbital of sulfur atoms have a binding energy of 239 MJ/mol.

# 20scoring Guidelines Ap Chemistry

**John Moore, Richard H. Langley**



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