

Use the activity series of elements to predict which reactions take place.

Metal	Metal Ions			
	Copper (Cu)	High (H)	Low (L)	Ag (Ag)
Copper (Cu)	<b>0%</b>	<b>100%</b> Max and Min	<b>100%</b> Max and Min	<b>100%</b>
Magnesium (Mg)	<b>100%</b>	<b>100%</b>	<b>100%</b>	<b>100%</b>
Zinc (Zn)	<b>100%</b>	<b>100%</b> Max and Min	<b>100%</b>	<b>100%</b>
Silver (Ag)	<b>100%</b> Max and Min	<b>100%</b> Max and Min	<b>100%</b> Max and Min	<b>100%</b>

### Procedure

1. Place your cells in corresponding numbered beakers/labeled beakers/labeled using this activity series as a guide.

Beaker 1: **Cu** **Ag**

2. Metals react with the most active metal: **Magnesium**

3. Metals react with the least active metal: **Ag**

4. Add the least reactive metal to each beaker. Observe the reaction. If the metal does not react with the most active metal, add the least active metal to the beaker. Observe the reaction. If the metal does not react with the least active metal, add the least active metal to the beaker. Observe the reaction. If the metal does not react with the least active metal, add the least active metal to the beaker. Observe the reaction.

Beaker 2: **Mg** **Cu** **Ag**

5. Metal to give metal activity test is observed. Observe the reaction. If the metal does not react with the most active metal, add the least active metal to the beaker. Observe the reaction.

**Copper is less reactive than Silver and is less active than Magnesium.**

6. Metal to give metal activity test is observed. Observe the reaction. If the metal does not react with the most active metal, add the least active metal to the beaker. Observe the reaction.

**Silver is less active than Copper and is less active than Magnesium.**

7. Metal to give metal activity test is observed. Observe the reaction. If the metal does not react with the most active metal, add the least active metal to the beaker. Observe the reaction.

# Activity Series Chemistry Lab Answers

**Ensheng Dong**



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