

		temperature:	change?
Ca	Should be described based on sample used in class	Vigorous bubbling, heat, white powdery forms on metal surface	Brilliant pink/fuchsia color
Mg	Should be described based on sample used in class	Little or no change, some small bubbles form on metal surface	A small bit of pink color
Al	Should be described based on sample used in class	No change	No change

Procedure

PART B: TO BE DEMONSTRATED BY THE TEACHER

1. The teacher will light the lab burner.
2. **Safety:** Do NOT stare directly at the Magnesium ribbon as you heat it. The teacher will add a pipette of water and 1 drop of phenolphthalein to the bottom of the Petri dish for the Mg. The teacher will hold the Mg over the small blue cone in the flame. Once it glows, the teacher will hold it OVER the Petri dish. **YOU WILL NOT BE ALLOWED TO LOOK DIRECTLY AT THE BURNING MAGNESIUM!** When finished burning, the teacher will place the Mg in the Petri dish. Record the color of the magnesium oxide and the color change in phenolphthalein in Data table 2.
3. The teacher will add a pipette of water and 1 drop of phenolphthalein to the bottom of the Petri dish for the Al.
4. The teacher will hold the Al over the small blue cone in the flame. Record observations of the reaction. When finished burning, the teacher will place the aluminum in the Petri dish. Record the color change with phenolphthalein in Data table 2.

Comparison of Magnesium to Aluminum

Data Table 2

Metal	Metal burning (how vigorously?)	Burned metal product + phenolphthalein
Ca	Calcium burns too vigorously to observe safely.	Using CaO if available, brilliant pink/fuchsia color
Mg	Very bright flame which SHOULD NOT be looked at directly	Some pink/fuchsia color
Al	Metal softens, glows brightly orange, hardens as soon as removed from flame	No change. Same result if Al_2O_3 is used

Critical Thinking: Answer each question using complete thoughts. Do not use the words "it" or "they" in your answers.

1. Elemental (unreacted) Ca and Mg metals are shiny when pure. What could have happened to the metals we used to make them become dull BEFORE we performed the lab activity?

Exposure to oxygen in the atmosphere produces a slower but identical result (oxidation)

2. Would you describe the reaction of Ca in water as exothermic (energy releasing) or endothermic (energy absorbing)? Use observations to support your answer.

Exothermic, petri dish becomes hot, steam is produced

3. Phenolphthalein appears pink in basic or alkaline solutions. Did the reactions of the alkaline earth metals with water produce alkaline solutions? Use observations to support your answer.

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Joacim Rocklöv



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1. Understanding the eBook Activity 35 Alkali Metal Clues Answers
 - The Rise of Digital Reading Activity 35 Alkali Metal Clues Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Activity 35 Alkali Metal Clues Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Activity 35 Alkali Metal Clues Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Activity 35 Alkali Metal Clues Answers

- Personalized Recommendations
- Activity 35 Alkali Metal Clues Answers User Reviews and Ratings
- Activity 35 Alkali Metal Clues Answers and Bestseller Lists
- 5. Accessing Activity 35 Alkali Metal Clues Answers Free and Paid eBooks
 - Activity 35 Alkali Metal Clues Answers Public Domain eBooks
 - Activity 35 Alkali Metal Clues Answers eBook Subscription Services
 - Activity 35 Alkali Metal Clues Answers Budget-Friendly Options
- 6. Navigating Activity 35 Alkali Metal Clues Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Activity 35 Alkali Metal Clues Answers Compatibility with Devices
 - Activity 35 Alkali Metal Clues Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Activity 35 Alkali Metal Clues Answers
 - Highlighting and Note-Taking Activity 35 Alkali Metal Clues Answers
 - Interactive Elements Activity 35 Alkali Metal Clues Answers
- 8. Staying Engaged with Activity 35 Alkali Metal Clues Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Activity 35 Alkali Metal Clues Answers
- 9. Balancing eBooks and Physical Books Activity 35 Alkali Metal Clues Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Activity 35 Alkali Metal Clues Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Activity 35 Alkali Metal Clues Answers
 - Setting Reading Goals Activity 35 Alkali Metal Clues Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Activity 35 Alkali Metal Clues Answers

- Fact-Checking eBook Content of Activity 35 Alkali Metal Clues Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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