

Common Polyatomic Ions

(Alphabetical order by ion name)

NOTE: -ite ending means one less oxygen than the -ate form.

Ion Name	Ion Symbol	Ion Name	Ion Symbol	Ion Name	Ion Symbol	Ion Name	Ion Symbol
Acetate	CH_3CO_2^- or CH_3COO^-	Dichromate	$\text{Cr}_2\text{O}_7^{2-}$	Hypobromite	BrO^-	Perrhenate	ReO_4^-
Ammonium	NH_4^+	Dihydrogen Phosphate	H_2PO_4^-	Hypochlorite	ClO^-	Phosphate	PO_4^{3-}
Arsenate	AsO_4^{3-}	Dihydrogen Phosphite	H_2PO_3^-	Hypoiodite	IO^-	Phosphite	PO_3^{3-}
Arsenite	AsO_3^{3-}	Dithionate	$\text{S}_2\text{O}_6^{2-}$	Iodate	IO_3^-	Plumbate	PbO_3^{2-}
Azide	N_3^-	Dithionite	$\text{S}_2\text{O}_4^{2-}$	Iodite	IO_2^-	Plumbite	PbO_2^{2-}
Borate	BO_3^{2-}	Ferricyanide	$\text{Fe}(\text{CN})_6^{3-}$	Isocyanate	NCO^-	Rhenate	ReO_4^{2-}
Bromate	BrO_3^-	Ferrocyanide	$\text{Fe}(\text{CN})_6^{4-}$	Manganate	MnO_4^{2-}	Selenate	SeO_4^{2-}
Bromite	BrO_2^-	Fulminate	CNO^-	Nitrate	NO_3^-	Selenite	SeO_3^{2-}
Carbonate	CO_3^{2-}	Hydrazide	N_2H_3^-	Nitrite	NO_2^-	Silicate	SiO_3^{2-}
Chlorate	ClO_3^-	Hydrogen arsenate	HAsO_4^{2-}	Oxalate	$\text{C}_2\text{O}_4^{2-}$	Stannate	SnO_3^{2-}
Chlorite	ClO_2^-	Hydrogen carbonate	HCO_3^-	Ozonide	O_3^-	Stannite	SnO_2^{2-}
Chromate	CrO_4^{2-}	Hydrogen Phosphate	HPO_4^{2-}	Perbromate	BrO_4^-	Sulfate	SO_4^{2-}
Chromite	CrO_2^-	Hydrogen Phosphite	HPO_3^{2-}	Perchlorate	ClO_4^-	Sulfite	SO_3^{2-}
Citrate	$\text{C}_6\text{H}_5\text{O}_7^{3-}$	Hydrogen Sulfate	HSO_4^-	Periodate	IO_4^-	Superoxide	O_2^-
Cyanate	OCN^-	Hydrogen Sulfite	HSO_3^-	Permanganate	MnO_4^-	Thiocyanate	SCN^-
Cyanide	CN^-	Hydroxide	OH^-	Peroxide	O_2^{2-}	Thiosulfate	$\text{S}_2\text{O}_3^{2-}$
						Tungstate	WO_4^{2-}

PREFIXES:

per- = one more oxygen than -ate

hypo- = one less oxygen than -ite

Activities For High School Chemistry Polyatomic Ions Packet Answers

R Bogdan



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