

Activity 1.3.3 Thermodynamics Answer Key

Introduction

Think back to the last time someone complained about a room being too warm. What did you notice about the temperature within the room? Is it really all that warm when? In Activity 1.3.3 you will investigate the effects of work, thermal energy, and energy in a system, all in the context of the water cycle that does take place.

Procedure

Answer the following questions as you finish the class the introduction to thermodynamics presentation.

1. Define thermodynamics.

The study of the transfer of work, thermal energy, and energy in a system.

2. List three examples of a thermodynamic system.

- a. **Boiling point**
- b. **Boiling water**
- c. **Boiling point**

3. Define thermal energy.

Thermal energy is heat that can be added to a system due to temperature difference.

4. Define temperature.

The average kinetic energy of particles in an object.

5. Fill in the table below with the correct temperature.

Boiler	Freezing point of water	Boiling point of water
Celsius	0°C	100°C
Fahrenheit	32°F	212°F
Kelvin	273K	373K

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