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web limiting reactant theoretical and percentage yields key $4\text{KO}_2(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 4\text{KOH}(\text{aq}) + 3\text{O}_2(\text{g})$ if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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web start with 8 3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8 3 \text{ mole } \text{H}_2\text{S} \times 2 \text{ mole } \text{H}_2\text{O} / 2 \text{ mole } \text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / 1 \text{ mole } \text{H}_2\text{O} = 149 4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $137 1 \text{ g } \text{H}_2\text{O} / 149 4 \text{ g } \text{H}_2\text{O} \times 100 = 91 77\%$ problems 6 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $160 0 \text{ g } \text{Al} \times 1 \text{ mol } \text{Al} / 27 0 \text{ g } \text{Al} \times 2 \text{ mol } \text{Al}_2\text{O}_3 / 4 \text{ mol } \text{Al} \times 101 964 \text{ g } \text{Al}_2\text{O}_3 / 302 3 \text{ g } \text{Al}_2\text{O}_3 = 26 982 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield $\text{actual} / \text{theoretical} \times 100 = 260 0 \text{ g } / 269 82 \text{ g } \times 100 = 96 01\%$

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