

## 8.2

# THE NATURE OF COVALENT BONDING

# KEY

## Section Review

### Objectives

- State a rule that usually tells how many electrons are shared to form a covalent bond
- Describe how electron dot formulas are used
- Predict when two atoms are likely to be joined by a double or a triple covalent bond
- Distinguish between a single covalent bond and other covalent bonds
- Describe how the strength of a covalent bond is related to its bond dissociation energy
- Describe how resonance structures explain bonding
- Identify some exceptions to the octet rule

### Vocabulary

- |                        |                            |                            |
|------------------------|----------------------------|----------------------------|
| • single covalent bond | • double covalent bonds    | • polyatomic ion           |
| • structural formulas  | • triple covalent bonds    | • bond dissociation energy |
| • unshared pairs       | • coordinate covalent bond | • resonance structures     |

### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

When atoms share electrons to gain the 1 configuration of a noble gas, the bonds formed are 2. A 3 pair of valence electrons constitutes a 4 covalent bond. Pairs of valence electrons that are not shared between atoms are called 5. Sometimes two or three pairs of electrons may be shared to give 6 covalent bonds. In some cases, only one of the atoms in a bond provides the pair of bonding electrons; this is a 7. 8 is required to break covalent bonds between atoms. The total energy required to break the bond between two covalently bonded atoms is known as the 9.

When it is possible to write two or more valid electron dot formulas for a molecule or ion, each formula is referred to as a 10.

- stable
- covalent
- shared
- single
- unshared pairs
- double/triple
- coordinate covalent bond
- energy
- bond dissociation energy
- resonance structure

## 8 2 The Nature Of Covalent Bonding Wkst Answers

**Phillip Manning**



## 8 2 The Nature Of Covalent Bonding Wkst Answers:

**The Nature of Matter Gr. 5-8 , Jacaranda Nature of Biology 2 VCE Units 3 and 4, LearnON and Print** Judith Kinnear, Marjory Martin, Lucy Cassar, Elise Meehan, Ritu Tyagi, 2021-10-29 Jacaranda Nature of Biology Victoria's most trusted VCE Biology online and print resource The Jacaranda Nature of Biology series has been rewritten for the VCE Biology Study Design 2022 2026 and offers a complete and balanced learning experience that prepares students for success in their assessments by building deep understanding in both Key Knowledge and Key Science Skills Prepare students for all forms of assessment Preparing students for both the SACs and exam with access to 1000s of past VCAA exam questions now in print and learnON new teacher only and practice SACs for every Area of Study and much more Videos by experienced teachers Students can hear another voice and perspective with 100s of new videos where expert VCE Biology teachers unpack concepts VCAA exam questions and sample problems For students of all ability levels All students can understand deeply and succeed in VCE with content mapped to Key Knowledge and Key Science Skills careful scaffolding and contemporary case studies that provide a real word context eLogbook and eWorkbook Free resources to support learning eWorkbook and the increased requirement for practical investigations eLogbook which includes over 80 practical investigations with teacher advice and risk assessments For teachers learnON includes additional teacher resources such as quarantined questions and answers curriculum grids and work programs

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**Biology/science Materials** Carolina Biological Supply Company, 1991

**SourceBook Version 2.1**, 1998

**Earth Science** Samuel N. Namowitz, Nancy E. Spaulding, 1989

**Biology**, 1987

**Resources for Teaching Middle School Science** Smithsonian Institution, National Academy of

Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-04-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them Resources for Teaching Middle School Science developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the National Science Education Standards This completely new guide follows on the success of Resources for Teaching Elementary School Science the first in the NSRC series of annotated guides to hands on inquiry centered curriculum materials and other resources for science teachers The curriculum materials in the new guide are grouped in five chapters by scientific area Physical Science Life Science Environmental Science Earth and Space Science and Multidisciplinary and Applied Science They are also grouped by type core materials supplementary units and science activity books Each annotation of curriculum material includes a recommended grade level a description of the activities involved and of what students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum materials included in this book were selected by panels of teachers and scientists using evaluation criteria developed for the guide The criteria reflect and incorporate goals and principles of the National Science Education Standards The annotations designate the specific content standards on which these curriculum pieces focus In addition to the curriculum chapters the guide contains six chapters of diverse resources that are directly relevant to middle school science Among these is a chapter on educational software and multimedia programs chapters on books about science and teaching directories and guides to science trade books and periodicals for teachers and students Another section features institutional resources One chapter lists about 600 science centers museums and zoos where teachers can take middle school students for interactive science experiences Another chapter describes nearly 140 professional associations and U S government agencies that offer resources and assistance Authoritative extensive and thoroughly indexed and the only guide of its kind Resources for Teaching Middle School Science will be the most used book on the shelf for science teachers school administrators teacher trainers science curriculum specialists advocates of hands on science teaching and concerned parents *Paperbound Books in Print*, 1975

**Chemical Bonds** Phillip Manning, 2009 Modern life is made up of a mind boggling array of materials A simple drinking cup for example might be made of Styrofoam paper or glass depending on the drinkers needs at the moment Home storage cabinets can be made of metal wood or plastic Space shuttles are assembled from silicon steel and hundreds of other materials All of these items owe their properties to the chemical bonds between the atoms that make up the substance Chemical Bonds examines the nature of the chemical bonds answering fundamental questions about how they form how they are broken and how they help define life as we know it *The Covalent Bond* Henry Sinclair Pickering, 1977 *The Chemical Bond* Gernot Frenking, Sason Shaik, 2014-06-13 This is the

perfect complement to *Chemical Bonding Across the Periodic Table* by the same editors who are two of the top scientists working on this topic each with extensive experience and important connections within the community The resulting book is a unique overview of the different approaches used for describing a chemical bond including molecular orbital based valence bond based ELF AIM and density functional based methods It takes into account the many developments that have taken place in the field over the past few decades due to the rapid advances in quantum chemical models and faster computers

**Chemistry of the Covalent Bond** Leallyn Burr Clapp,1957      **How Chemical Bonds Form and Chemical Reactions Proceed** V. I. Gankin, Yuri V. Gankin,1998      The Nature of Ultralong C-C Bonds Takuya Shimajiri,2023 This book focuses on the creation and demonstration of a carbon carbon C C single bond beyond 1.8 and on elucidation of its unique nature C C single bond is one of the most fundamental concepts in organic chemistry The elucidation of its nature is important for further understanding chemical phenomena The nature in the extreme state of C C single bond is still unexplored because of the instability of compounds In terms of its bond length the limit had been predicted around 1.8 based on the experimental and theoretical studies This book describes a first example of a C C single bond beyond 1.8 by employing the original intramolecular core shell strategy to make a weak and elongated bond stable enough The presence of such an elongated bond was demonstrated by experimental and theoretical studies The bond length changes could affect physical properties such as optical absorption and redox potential Furthermore its unique flexibility was discovered This book benefits the chemists with deeper understanding of the covalent bonding      *The Nature of the Chemical Bond* Linus Pauling,1960 Resonance and the chemical bond The partial ionic character of covalent bonds and the relative electronegativity of atoms The directed bond Bond strengths and bond angles The magnetic criterion for bond type The resonance of molecules among several valence bond structures Interatomic distances and their relation to the structure of molecules and crystals Types of resonance in molecules The structure of molecules and complex ions involving bonds with partial double bond character The one electron bond and the three electron bond The hydrogen bond The sizes of ions and the structure of ionic crystals The metallic bond A summarizing discussion of resonance and its significance for chemistry      Applications of Reversible Covalent Bonding Within the Context of Spin Trapping, Boronate Ester Formation, Conjugate Addition of Amines and Thiols, and the Synthesis of Natural Product Analogues Margaret Katherine Meadows,2016 Reversible covalent bonding is an essential chemical tool that incorporates the strength of covalent bonds and the dynamism of intermolecular interactions This type of bonding has found applications across fields of chemistry including sensing of small molecules screening against target for the design of host guest systems and enzyme inhibition Herein are discussed examples of reversible covalent bonding A three component assembly incorporating 2 formylphenyl boronic acid N alkylhydroxylamine and catechol was synthesized with the intention of acting as a reversible spin trap for nitric oxide detection This assembly was further studied as a potential bioorthogonal method for specific chemical modification of amino acids within a peptide or protein This system

is amenable to a variety of substrates including L DOPA a catechol containing amino acid which can readily be incorporated into proteins This assembly is also compatible with existing bioorthogonal reactions allowing for multiple labeling reactions to be conducted simultaneously The ability to reversibly crosslink two substrates under mild conditions has the potential for wide application Our group has previously developed a conjugate acceptor derived from Meldrum's acid that sequentially crosslinks an amine and a thiol under aqueous conditions at neutral pH This crosslinking can be decoupled with a chemical trigger to regenerate the original species We have since expanded this system to include the addition of aniline derivatives We have further optimized this conjugate acceptor and studied the pK<sub>a</sub>'s and reaction rates of the decoupling using different substrates The natural product piperlongumine has been found to induce apoptosis in cancer cells with little effect on normal cells and is thought to act through increasing the level of reactive oxygen species ROS This increase in ROS is proposed to occur through the reversible addition of glutathione into the more reactive Michael acceptor followed by the addition of a glutathione binding protein to the second Michael acceptor We have synthesized piperlongumine using a route that can be easily modified to access derivatives with various substituents in order to study the potency of analogs and to test the proposed mechanism

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