

KEY26. Which is the stronger base, HCO_3^- or HCOO^- ? (1 mark) HCO_3^-

27. Classify each of the following as: a strong acid (SA), weak acid (WA), strong base (SB), weak base (WB) or a spectator ion (S). (10 marks)

a) F^-	WB	f) Cl^-	S
b) HIO_3	WA	g) NH_3	WB
c) NO_3^-	S	h) O^{2-}	SB
d) HClO_4	SA	i) CH_3COOH	WA
e) $\text{C}_2\text{O}_4^{2-}$	WB	j) ClO_4^-	S

28. What is the $[\text{OH}^-]$ in a solution made by adding 0.060 moles of calcium oxide to 500.0 mL water? Be careful! (2 marks)

$$M = \frac{\text{mol}}{L} = \frac{0.060 \text{ mol CaO}}{0.5000 \text{ L}} = 0.12 \text{ M} = [\text{CaO}]$$

$$0.12 \text{ CaO} \rightarrow \text{Ca}^{2+} + 0.12 \text{ O}^{2-} \Rightarrow \text{O}^{2-} + \text{H}_2\text{O} \rightarrow 2\text{OH}^-$$

Answer

 $[\text{OH}^-] = 0.24 \text{ M}$ 29. What is the $[\text{H}_3\text{O}^+]$ in a solution made by adding 0.020 moles of nitric acid to 500.0 mL of water? (2 marks)

$$[\text{HNO}_3] = \frac{0.020 \text{ mol}}{0.500 \text{ L}} = 0.040 \text{ M} \quad \text{SA so } [\text{H}_3\text{O}^+] = [\text{acid}] = 0.040 \text{ M}$$

Answer

 $[\text{H}_3\text{O}^+] = 0.040 \text{ M}$ 30. If 0.10 M HSO_3^- is mixed with 0.10 M HC_2O_4^- , which species will donate a proton? (1 mark)

Answer

 HC_2O_4^- 31. If 0.10 M HSO_3^- is mixed with 0.10 M $\text{HC}_2\text{H}_3\text{O}_2$, which species will donate a proton? (1 mark)

Answer

 HSO_3^-

19 3 Acids And Bases Answer Key Chemistry

Simeon Shaw



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