

Name: \_\_\_\_\_ per: \_\_\_\_\_

# Worksheet - Reaction Rates

Use this reaction for the questions below:  $C_6H_{12}O_6 (s) + 6 O_2(g) \rightarrow 6 H_2O (g) + 6 CO_2 (g)$

- What happens to the concentrations of:
  - $C_6H_{12}O_6$  &  $O_2$  as the reaction proceeds  $\rightarrow$ ?
  - $H_2O$  +  $CO_2$  as the reaction proceeds  $\rightarrow$ ?
- According to the collision theory, what 3 circumstances are needed for  $C_6H_{12}O_6$  &  $O_2$  to react?
- What is the activation energy for a chemical reaction?
- Use the equation & the collision theory to explain:



Change in condition:	Does this <u>increase</u> or <u>decrease</u> the rate of reaction?	Explain why
a. <u>Increasing the temperature</u>	Ex: Increase (springs up)	Ex: Molecules move faster & collide more = increased rxn rate
b. <u>Increasing the concentration</u> of $C_6H_{12}O_6$		
c. <u>Decreasing the concentration</u> of $O_2$		
d. <u>Increase the surface area</u> by chewing up food in your mouth.		
e. <u>Decreasing the temperature</u>		
f. <u>Increasing the pressure</u> in the container		
g. <u>Decreasing the concentration</u> of $H_2O$		
h. <u>Increasing the volume</u> of the container the reaction occurs in		
i. <u>Increasing the concentration</u> of $CO_2$		
j. <u>Using a catalyst</u> (like salivary amylase)		

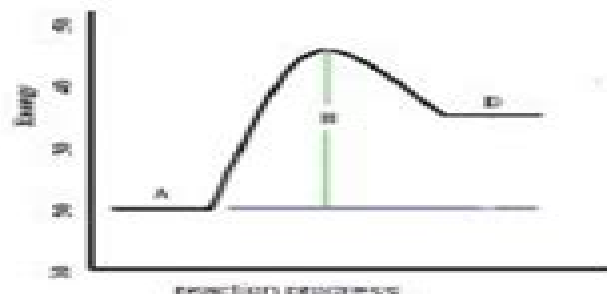
- On the accompanying energy diagram, label the following terms:

a. reactants    b. products    c. activation energy

- On the graph to the right, draw and label what this diagram would look like if a catalyst was added to the reaction.

- Graph reading

- How much energy (kJ) do the reactants have?
- How much energy (kJ) do the products have?
- How much energy (kJ) is required to activate this un-catalyzed reaction?
- Is this reaction endothermic or exothermic? How do you know?



# 18 Reaction Rates And Equilibrium Worksheet Answers

**Steven Jobe Onysko**



## 18 Reaction Rates And Equilibrium Worksheet Answers:

Software for Teaching Science Roger Frost,1998      Chemical Kinetics with Mathcad and Maple Viktor Korobov,Valery Ochkov,2011-05-26 The authors explain at length the principles of chemical kinetics and approaches to computerized calculations in modern software suites mathcad and maple Mathematics is crucial in determining correlations in chemical processes and requires various numerical approaches Often significant issues with mathematical formalizations of chemical problems arise and many kinetic problems can't be solved without computers Numerous problems encountered in solving kinetics calculations with detailed descriptions of the numerical tools are given Special attention is given to electrochemical reactions which fills a gap in existing texts not covering this topic in detail The material demonstrates how these suites provide quick and precise behavior predictions for a system over time for postulated mechanisms Examples i e oscillating and non isothermal reactions help explain the use of mathcad more efficiently Also included are the results of authors own research toward effective computations      **Books in Print Supplement** ,2002      *Radioactive Waste Management* ,1981

**Chemical Abatement of Acid Mine Drainage Formation** Steven Jobe Onysko,1986      **Report of the Annual Meeting of the South African Association for the Advancement of Science** ,2005      *INIS Atomindex* ,1984

Government Reports Announcements & Index ,1987-05      **Workbook 15** Thobeka Pete,Turret Correspondence College (Johannesburg),1988      **CHEMICAL EQUILIBRIUM** NARAYAN CHANGDER,2024-04-01 Note Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u gmail com You can also get full PDF books in quiz format on our youtube channel <https://www.youtube.com/smartquiziz> I will send you a PDF version of this workbook This book has been designed for candidates preparing for various competitive examinations It contains many objective questions specifically designed for different exams Answer keys are provided at the end of each page It will undoubtedly serve as the best preparation material for aspirants This book is an engaging quiz eBook for all and offers something for everyone This book will satisfy the curiosity of most students while also challenging their trivia skills and introducing them to new information Use this invaluable book to test your subject matter expertise Multiple choice exams are a common assessment method that all prospective candidates must be familiar with in today's academic environment Although the majority of students are accustomed to this MCQ format many are not well versed in it To achieve success in MCQ tests quizzes and trivia challenges one requires test taking techniques and skills in addition to subject knowledge It also provides you with the skills and information you need to achieve a good score in challenging tests or competitive examinations Whether you have studied the subject on your own read for pleasure or completed coursework it will assess your knowledge and prepare you for competitive exams quizzes trivia and more      **Theories of Chemical Reaction Rates** Keith James Laidler,1971      **The Equilibrium Assumption in the Theory of Absolute Reaction Rates** Charles Francis Curtiss,1948 The assumption that systems in the activated state are in equilibrium with the reactant molecules has been open to question In this paper the

methods of classical mechanics are used to obtain an expression for the rate constant without making this assumption. The result is similar in form to that obtained by Eyring except that the average velocity of the points in configuration space is related directly to the distribution in velocities of the molecules of the gas in real space. If one assumes that this distribution is Maxwellian, one obtains the Eyring expression. However, the chemical reaction itself acts as a perturbing influence and causes the distribution in velocities to be somewhat different from Maxwellian. It is this perturbation of the distribution function which leads to a different value of the rate constant and consequently to a measure of the effect of the equilibrium assumption.

Author      **Rates of Reaction and Equilibria** Clifford Othen, 1972

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## **Table of Contents 18 Reaction Rates And Equilibrium Worksheet Answers**

1. Understanding the eBook 18 Reaction Rates And Equilibrium Worksheet Answers
  - The Rise of Digital Reading 18 Reaction Rates And Equilibrium Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying 18 Reaction Rates And Equilibrium Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 18 Reaction Rates And Equilibrium Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from 18 Reaction Rates And Equilibrium Worksheet Answers
  - Personalized Recommendations
  - 18 Reaction Rates And Equilibrium Worksheet Answers User Reviews and Ratings
  - 18 Reaction Rates And Equilibrium Worksheet Answers and Bestseller Lists
5. Accessing 18 Reaction Rates And Equilibrium Worksheet Answers Free and Paid eBooks
  - 18 Reaction Rates And Equilibrium Worksheet Answers Public Domain eBooks
  - 18 Reaction Rates And Equilibrium Worksheet Answers eBook Subscription Services
  - 18 Reaction Rates And Equilibrium Worksheet Answers Budget-Friendly Options
6. Navigating 18 Reaction Rates And Equilibrium Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - 18 Reaction Rates And Equilibrium Worksheet Answers Compatibility with Devices
  - 18 Reaction Rates And Equilibrium Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 18 Reaction Rates And Equilibrium Worksheet Answers
  - Highlighting and Note-Taking 18 Reaction Rates And Equilibrium Worksheet Answers
  - Interactive Elements 18 Reaction Rates And Equilibrium Worksheet Answers

8. Staying Engaged with 18 Reaction Rates And Equilibrium Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 18 Reaction Rates And Equilibrium Worksheet Answers
9. Balancing eBooks and Physical Books 18 Reaction Rates And Equilibrium Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 18 Reaction Rates And Equilibrium Worksheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 18 Reaction Rates And Equilibrium Worksheet Answers
  - Setting Reading Goals 18 Reaction Rates And Equilibrium Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 18 Reaction Rates And Equilibrium Worksheet Answers
  - Fact-Checking eBook Content of 18 Reaction Rates And Equilibrium Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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