

Reversible Reactions & Equilibrium

Question 1

Hydrated cobalt(II) chloride undergoes thermal decomposition when heated to produce anhydrous cobalt(II) chloride and water.



The addition of water to CoCl_2 rehydrates the compound. Which row correctly describes the reverse reaction?

	reaction type	colour change
A	exothermic	blue to pink
B	endothermic	blue to pink
C	exothermic	pink to blue
D	endothermic	pink to blue

[1 mark]

Question 2

Which of the following reactions is reversible?

- A. $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$
- B. $\text{CuSO}_4 \cdot 5\text{H}_2\text{O} \rightarrow \text{CuSO}_4 + 5\text{H}_2\text{O}$
- C. $\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{FeSO}_4 + \text{H}_2$
- D. $\text{KCl} + \text{AgNO}_3 \rightarrow \text{KNO}_3 + \text{AgCl}$

[1 mark]

Question 3

Carbon and steam react together as shown in the equation.



Both of the products from the forward reaction are gaseous. Which conditions of pressure and temperature would provide the highest yield of H_2 ?

	pressure	temperature
A	low	low
B	low	high
C	high	low
D	high	high

[1 mark]

18 2 Reversible Reactions And Equilibrium Answer Key

L Darling-Hammond



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