

EXPERIMENT: INVESTIGATING PRECIPITATION REACTIONS

Syllabus reference 8.4.3

AIM

To compare the solubility of a range of ionic substances through precipitation reactions.

EQUIPMENT

- test tubes (6)

Cations (all 0.1 mol L^{-1}) in dropper bottles

- sodium nitrate, NaNO_3
- silver nitrate, AgNO_3
- lead(II) nitrate, $\text{Pb}(\text{NO}_3)_2$
- ammonium nitrate, NH_4NO_3
- magnesium sulfate, MgSO_4
- iron(II) sulfate, FeSO_4
- copper(II) sulfate, CuSO_4
- zinc(II) sulfate, ZnSO_4
- aluminium sulfate, $\text{Al}_2(\text{SO}_4)_3$
- barium chloride, BaCl_2
- calcium chloride, CaCl_2
- iron(III) chloride, FeCl_3

- test tube rack

Anions (all 0.1 mol L^{-1}) in dropper bottles

- potassium nitrate, KNO_3
- potassium chloride, KCl
- sodium sulfate, Na_2SO_4
- sodium carbonate, Na_2CO_3
- sodium acetate, NaCH_3COO
- sodium hydroxide, NaOH



SAFETY: Wear safety glasses. Lead and barium salts are poisonous so avoid contact with skin. If contact occurs wash thoroughly with soap and water. Dispose of as directed by your teacher. Consult 'Chemical Safety in Schools' booklet (CSIS).

PROCEDURE

- 1 Put a 5 mL of a chemical containing each anion into each of six test tubes. Take care not to mix droppers.
- 2 Add 5 mL of a chemical containing the first cation to each. Record your observations.
- 3 Clean all test tubes thoroughly as incorrect results can arise from contamination of solutions.

17 Precipitation Reactions Lab Worksheet

Firman Edward Bear



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Table of Contents 17 Precipitation Reactions Lab Worksheet

1. Understanding the eBook 17 Precipitation Reactions Lab Worksheet
 - The Rise of Digital Reading 17 Precipitation Reactions Lab Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying 17 Precipitation Reactions Lab Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 17 Precipitation Reactions Lab Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from 17 Precipitation Reactions Lab Worksheet
 - Personalized Recommendations
 - 17 Precipitation Reactions Lab Worksheet User Reviews and Ratings

- 17 Precipitation Reactions Lab Worksheet and Bestseller Lists
- 5. Accessing 17 Precipitation Reactions Lab Worksheet Free and Paid eBooks
 - 17 Precipitation Reactions Lab Worksheet Public Domain eBooks
 - 17 Precipitation Reactions Lab Worksheet eBook Subscription Services
 - 17 Precipitation Reactions Lab Worksheet Budget-Friendly Options
- 6. Navigating 17 Precipitation Reactions Lab Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - 17 Precipitation Reactions Lab Worksheet Compatibility with Devices
 - 17 Precipitation Reactions Lab Worksheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 17 Precipitation Reactions Lab Worksheet
 - Highlighting and Note-Taking 17 Precipitation Reactions Lab Worksheet
 - Interactive Elements 17 Precipitation Reactions Lab Worksheet
- 8. Staying Engaged with 17 Precipitation Reactions Lab Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 17 Precipitation Reactions Lab Worksheet
- 9. Balancing eBooks and Physical Books 17 Precipitation Reactions Lab Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 17 Precipitation Reactions Lab Worksheet
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 17 Precipitation Reactions Lab Worksheet
 - Setting Reading Goals 17 Precipitation Reactions Lab Worksheet
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 17 Precipitation Reactions Lab Worksheet
 - Fact-Checking eBook Content of 17 Precipitation Reactions Lab Worksheet
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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