

Polyatomic Ions Worksheet Answer Key

Polyatomic Practice

1. Name or write the formula for the following polyatomic ions

sulfate	<u>SO_4^{2-}</u>	<u>CO_3^{2-}</u>	carbonate
nitrite	<u>NO_2^-</u>	<u>MnO_4^{1-}</u>	manganate
perphosphate	<u>PO_4^{3-}</u>	<u>SO_5^{2-}</u>	persulfate
hypoiodite	<u>IO^-</u>	<u>BrO_2^{1-}</u>	bromite
chlorite	<u>ClO_2^-</u>	<u>CO_4^{2-}</u>	percarbonate
phosphite	<u>PO_3^{3-}</u>	<u>PO_5^{3-}</u>	perphosphate
percarbonate	<u>CO_4^{2-}</u>	<u>ClO^{1-}</u>	hypochlorite
bromate	<u>BrO_3^-</u>	<u>IO_2^{1-}</u>	iodite
hyposulfite	<u>SO_2^{2-}</u>	<u>PO_4^{3-}</u>	phosphate
permanganate	<u>MnO_4^-</u>	<u>NO_2^{1-}</u>	nitrite
carbonite	<u>CO_2^{2-}</u>	<u>SO_4^{2-}</u>	sulfate

2. Name or write the formula for the following Type I polyatomic ionic compounds

beryllium hydroxide	<u>$\text{Be}(\text{OH})_2$</u>	<u>$\text{Ba}(\text{IO}_3)_2$</u>	barium iodate
sodium nitrite	<u>NaNO_2</u>	<u>$\text{Ga}(\text{CNO})_3$</u>	gallium cyanate
ammonium chloride	<u>NH_4Cl</u>	<u>Ag_2SO_3</u>	silver sulfite
calcium bisulfate	<u>$\text{Ca}(\text{HSO}_4)_2$</u>	<u>MgCO</u>	magnesium hypocarbonite
rubidium perchlorate	<u>RbClO_4</u>	<u>NH_3NO_2</u>	ammonium nitrite
strontium sulfite	<u>SrSO_3</u>	<u>$\text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3$</u>	aluminum acetate
aluminum acetate	<u>$\text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3$</u>	<u>SrSO_5</u>	strontium persulfate
ammonium nitrate	<u>NH_4NO_3</u>	<u>RbClO_2</u>	rubidium chlorite
magnesium hypocarbonite	<u>MgCO</u>	<u>$\text{Ca}(\text{HSO}_4)_2$</u>	calcium bisulfate
silver hyposulfite	<u>Ag_2SO_2</u>	<u>NH_4Cl</u>	ammonium chloride

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Matthew Johll



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