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Chemical Periodicity

1. Which of the following orders of ionic radii is correctly represented? [2014]

- (a) $H^- > H^+ > H$ (b) $Na^+ > F^- > O^{2-}$
(c) $F^- > O^{2-} > Na^+$ (d) $Al^{3+} > Mg^{2+} > N^{3-}$

2. Identify the wrong statement in the following. [2012]

- (a) Amongst isoelectronic species, smaller the positive charge on the cation, smaller is the ionic radius
(b) Amongst isoelectronic species, greater the negative charge on the anion, larger is the ionic radius
(c) Atomic radius of the elements increases as one moves down the first group of the periodic table
(d) Atomic radius of the elements decreases as one moves across from left to right in the 2nd period of the periodic table

3. The correct order of the decreasing ionic radii among the following isoelectronic species is [2010]

- (a) $Ca^{2+} > K^+ > S^{2-} > Cl^-$
(b) $Cl^- > S^{2-} > Ca^{2+} > K^+$
(c) $S^{2-} > Cl^- > K^+ > Ca^{2+}$
(d) $K^+ > Ca^{2+} > Cl^- > S^{2-}$

4. Which of the following represents the correct order of increasing electron gain enthalpy with negative sign for the elements O, S, F and Cl? [2010]

- (a) $Cl < F < O < S$ (b) $O < S < F < Cl$
(c) $F < S < O < Cl$ (d) $S < O < Cl < F$

5. Amongst the elements with following electronic configurations, which one may have the highest ionisation energy? [2009]

- (a) $[Ne] 3s^2 3p^2$
(b) $[Ne] 3s^2 3p^3$
(c) $[Ar] 3d^{10}, 4s^2 4p^3$
(d) $[Ne] 3s^2 3p^3$

6. Which one of the elements with the following outer orbital configurations may exhibit the largest number of oxidation states? [2009]

- (a) $3d^3, 4s^2$ (b) $3d^5, 4s^1$ (c) $3d^5, 4s^2$ (d) $3d^2, 4s^2$

7. The correct order of decreasing second ionisation enthalpy of Ti(22), V(23), Cr(24) and Mn(25) is [2008]

- (a) $Cr > Mn > V > Ti$ (b) $V > Mn > Cr > Ti$
(c) $Mn > Cr > Ti > V$ (d) $Ti > V > Cr > Mn$

8. Which of the following electronic configuration of an atom has the lowest ionisation enthalpy? [2007]

- (a) $1s^2, 2s^2 2p^5$ (b) $1s^2, 2s^2 2p^3$
(c) $1s^2, 2s^2 2p^6, 3s^1$ (d) $1s^2, 2s^2 2p^6$

9. Identify the correct order of the size of the following. [2007]

- (a) $Ca^{2+} < K^+ < Ar < S^{2-} < Cl^-$
(b) $Ca^{2+} < K^+ < Ar < Cl^- < S^{2-}$
(c) $Ar < Ca^{2+} < K^+ < Cl^- < S^{2-}$
(d) $Ca^{2+} < Ar < K^+ < Cl^- < S^{2-}$

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Ted Lister, Janet Renshaw



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