

14-1 Review and Reinforcement

Condensed States of Matter

Fill in the following chart comparing solids and liquids.

Physical Property	Liquids	Solids
1. Compressibility		
2. Density		
3. Shape		
4. Rate of Diffusion		

Answer each of the following questions in the space provided.

5. Describe the behavior of the molecules in a liquid. Explain this behavior in terms of intermolecular forces.
- _____
- _____
6. How do intermolecular forces determine molecular arrangement in solids?
- _____
- _____

If the statement is true, write "true." If it is false, change the underlined word or words to make the statement true. Write your answer on the line provided.

- _____ 7. Intermolecular forces determine metallic properties such as the boiling point of a substance.
- _____ 8. Intermolecular forces are forces of attraction between atoms.
- _____ 9. Chemical bonds are the forces between molecules.
- _____ 10. Covalent bonds result from electrons being shared between atoms in a molecule.
- _____ 11. Intermolecular forces result from the electron interactions between neighboring molecules.
- _____ 12. When the electron cloud is not distributed symmetrically, a molecule has a polarity.
- _____ 13. Chemical bonds include ionic bonds, metallic bonds, and atomic bonds.
- _____ 14. Intermolecular forces include dispersion forces, dipole-dipole forces, and helium-bond forces.

14 Review And Reinforcement Answers

Emmett Sergio Walls



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