

The Gas Laws – Ch. 10 (p.313-322)

1. The gas left in a used aerosol can is at a pressure of 1 atm at 27°C. If this can is thrown into a fire, what is the internal pressure of the gas when its temperature reaches 927°C?

GIVEN	GAS LAW	WORK
T1= 300K P1= 101 kPa T2= 1200K P2= x	BOYLE	101 / 300 = x / 1200 x = 404
	FORMULA	
	P1/T1=P2/T2	
ANSWER: 404 kPa OR 4 ATM		

2. A sample of carbon dioxide occupies a volume of 3.50 L at 125 kPa. What pressure would the gas exert if the volume were decreased to 2.00 L?

GIVEN	GAS LAW	WORK
V1 = 3.50 L P1 = 125 kPa V2 = 2.00 L P2 = x	CHARLES	(125) (3.50) = (2.00) (x) x = 218.75 kPa
	FORMULA	
	P1V1 = P2V2	
ANSWER: 218.75 kPa		

3. A sample of propane occupies 250.0 L at 125 kPa and 38°C. Find its volume at 100.0 kPa and 95°C.

GIVEN	GAS LAW	WORK
v1 = 250.0 L P1 = 125 kPa T1 = 311 K V2 = x P2 = 100.0 kPa T2 = 368 K	COMBINED	(125) (250.0) / 311 = (100.0) (x) / 368 x = 369.77
	FORMULA	
	P1v1/T1=P2v2/T2	
ANSWER: 369.77 L		

142 The Gas Laws Worksheet Answers

N Colangelo

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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