

Time : 2 Hrs.

Max. Marks : 40

- Q. 1** Select and write the most appropriate answer from the given alternatives for each sub-question. [8]
- (i) The addition of the non-volatile solute into the pure solvent ..... (1)  
 (a) Increases the vapour pressure of solvent. (b) Decreases the boiling point of solvent.  
 (c) Decreases the freezing point of solvent. (d) Increases the freezing point of solvent.
- (ii) What is true for an adiabatic process ? (1)  
 (a)  $\Delta T = 0$  (b)  $\Delta E = 0$  (c)  $\Delta H = \Delta E$  (d)  $q = 0$
- (iii) An aqueous solution of which one of the following salts is basic ? (1)  
 (a)  $\text{CH}_3\text{COONa}$  (b)  $\text{NH}_4\text{Cl}$  (c)  $\text{KNO}_3$  (d)  $\text{CuSO}_4$
- (iv) Which one of the following is the best projectile ? (1)  
 (a) Proton (b) Neutron (c)  $\alpha$ -particle (d) Deuteron
- (v) In a bimolecular reaction if one of the reactant is in large excess, then order of reaction is said to be ..... (1)  
 (a) Zero (b) Pseudo First (c) Second (d) Third
- (vi) Which one of the following transition elements show highest oxidation state ? (1)  
 (a) Sc (b) Ti (c) Mn (d) Zn
- (vii) The electromotive force of the following cell : (1)  
 $\text{Cu} | \text{Cu}^{++} (1\text{M}) || \text{Ag}^+ (1\text{M}) | \text{Ag}$  is ..... if  $E_{\text{Cu}^{++}/\text{Cu}}^0 = 0.33 \text{ V}$  and  $E_{\text{Ag}^+/\text{Ag}}^0 = 0.79 \text{ V}$   
 (a) 0.46 V (b) -0.46 V (c) 1.12 V (d) -1.12 V
- (viii) A gas expands from a volume of  $5 \times 10^{-3} \text{ m}^3$  to  $20 \times 10^{-3} \text{ m}^3$  at a pressure of  $2.5 \times 10^5 \text{ Nm}^{-2}$ . The work done is ..... (1)  
 (a)  $37.5 \times 10^2 \text{ J}$  (b)  $3.75 \times 10^2 \text{ J}$  (c)  $-3.75 \times 10^2 \text{ J}$  (d)  $-37.5 \times 10^2 \text{ J}$
- Q. 2 (A)** Attempt any One : [8]
- (i) Distinguish between molecularity and order of reaction. (2)  
 (ii) Derive the relationship between heat of a reaction at constant pressure and at constant volume. (2)
- (B)** Attempt any One : (2)  
 (i) Explain Lewis theory of acids and bases. (2)  
 (ii) Define : (a) Colligative Properties, (b) Solubility Product. (2)
- (C)** Answer the following : (2)  
 (i) Explain the position of d-block elements in the periodic table. (2)  
 (ii) Mention different ways of expressing concentration of a solution. (2)
- Q. 3 (A)** Attempt any One : [8]
- (i) What is the action of the following on zinc ? (3)  
 (a) dil.  $\text{H}_2\text{SO}_4$ , (b) Hot and Conc.  $\text{NaOH}$ , (c)  $\text{CuSO}_4$  solution.
- (ii) What is mass defect and binding energy ? Give the mathematical expression for binding energy per nucleon. (3)
- (B)** Attempt any One : (3)  
 (i) Write a note on de-electronation and electronation. (3)  
 (ii) Define rate of reaction. Derive integrated rate equation for first order reaction. (3)
- (C)** Answer the following :  
 Show that the degree of dissociation of weak acid is inversely proportional to square root of concentration. (2)
- Q. 4 (A)** Answer the following : [8]
- State and explain Hess's law of constant heat summation. Give its any four applications. (4)
- (B)** Attempt any One : (4)  
 (i) Describe the construction and working of Standard Hydrogen Electrode. Give any two disadvantages of it. (4)  
 (ii) Describe Ostwald-Walker method for determination of molecular weight by lowering of vapour pressure. (4)
- Q. 5 (A)** Attempt any One : [8]
- (i) Heat of following of  $\text{NH}_3(g)$  at 300 K is  $-46.2 \text{ kJ mol}^{-1}$ . Calculate its value at (a) 400 K and (b) 250 K. The mean molar heat capacities of  $\text{N}_2(g)$ ,  $\text{H}_2(g)$  and  $\text{NH}_3(g)$  at constant pressure and in the given range of temperature are 28.4, 28.3 and  $34.5 \text{ JK}^{-1}$  respectively. (4)

# 2008 Hsc Chemistry Answers

**2018 Edit, Mark Dixon**



## 2008 Hsc Chemistry Answers:

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### **Table of Contents 2008 Hsc Chemistry Answers**

1. Understanding the eBook 2008 Hsc Chemistry Answers
  - The Rise of Digital Reading 2008 Hsc Chemistry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying 2008 Hsc Chemistry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in a 2008 Hsc Chemistry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2008 Hsc Chemistry Answers
  - Personalized Recommendations
  - 2008 Hsc Chemistry Answers User Reviews and Ratings
  - 2008 Hsc Chemistry Answers and Bestseller Lists

5. Accessing 2008 Hsc Chemistry Answers Free and Paid eBooks
  - 2008 Hsc Chemistry Answers Public Domain eBooks
  - 2008 Hsc Chemistry Answers eBook Subscription Services
  - 2008 Hsc Chemistry Answers Budget-Friendly Options
6. Navigating 2008 Hsc Chemistry Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - 2008 Hsc Chemistry Answers Compatibility with Devices
  - 2008 Hsc Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 2008 Hsc Chemistry Answers
  - Highlighting and Note-Taking 2008 Hsc Chemistry Answers
  - Interactive Elements 2008 Hsc Chemistry Answers
8. Staying Engaged with 2008 Hsc Chemistry Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 2008 Hsc Chemistry Answers
9. Balancing eBooks and Physical Books 2008 Hsc Chemistry Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 2008 Hsc Chemistry Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 2008 Hsc Chemistry Answers
  - Setting Reading Goals 2008 Hsc Chemistry Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2008 Hsc Chemistry Answers
  - Fact-Checking eBook Content of 2008 Hsc Chemistry Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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