	In the following questions one mark each (Q. No. 6 - 9) a statement of **Assertion (A)** followed by a statement of **Reason (R)** is given. Choose the correct answer out of the following choices. (i) A and R both are correct statements and R is the correct explanation for A. (ii) A and R both are correct statements and R is not correct explanation for A. (iii) A is correct statement but R is wrong statement. (iv) A is wrong statement but R is correct statement.	
6.	**Assertion (A):** The empirical mass of ethene is half of its molecular mass. **Reason (R):** The empirical formula represents the simplest whole-number ratio of various atoms present in a compound.	1
7.	**Assertion (A):** Equal moles of different substances contain same number of constituent particles. **Reason (R):** Equal weights of different substances contain the same number of constituent particles.	1
8.	**Assertion:** It is impossible to determine the exact position and exact momentum of an electron simultaneously. **Reason:** The path of an electron in an atom is clearly defined.	1
9.	**Assertion:** The radius of the first orbit of hydrogen atom is 0.529Å. **Reason:** Radius of each circular orbit (r_n) = 0.529Å (n^2/Z), where n = 1, 2, 3 and Z = atomic number.	1
10.	45.4 L of dinitrogen reacted with 22.7 L of dioxygen and 45. 2/4 nitrous oxide was formed. The reaction is given below: $2N_2(g) + O_2(g) \rightarrow 2N_2O(g)$ Which law is being obeyed in this experiment? Write the statement of the law.	
11.	The density of the 3 molal solution of NaOH is 1.110 g mL^{-1}. Calculate the molarity of the solution.	2
12.	What is the difference between the terms orbit and orbital?	2
13	A cricket ball of mass 100 g does not move like a wave when it is thrown by a bowler at a speed of 100 km/h. Calculate the wavelength of the ball and explain why it does not show wave nature.	2
14.	Calculate the mass percent of calcium, phosphorus and oxygen in calcium phosphate $Ca_3(PO_4)_2$	3
15.	(a) 45.4 L of dinitrogen reacted with 22.7 L of dioxygen and 45.4 L of nitrous oxide was formed. The reaction is given below: $2N_2(g) + O_2(g) \rightarrow 2N_2O(g)$, Which law is being obeyed in this experiment? Write the statement of the law. (b) How many significant figures should be present in answer to the following calculation? $\frac{2.5 \times 1.25 \times 3.5}{2.01}$	3

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