

Time allowed: 2:40 Hours

Total Marks Sections B and C:

NOTE: Answer any eleven parts from Section 'B' and any two questions from Section 'C' on the separate provided answer book. Write your answers neatly and legibly. 09280208

SECTION - B (Marks 33)

(11 × 3 =

Q. 2 Attempt any ELEVEN parts. All parts carry equal marks.

- (i) Briefly explain the following. Also give one example of each
 - (a) Ion
 - (b) Environmental Chemistry
 - (c) Atomic Mass
- (ii) Compare compounds and mixtures. Also give examples.
- (iii) There are three isotopes of hydrogen. Draw their diagrams.
- (iv) Distinguish between 'shell' and 'subshell'.
- (v) State periodic law. Also write the names of groups in s block of the periodic table.
- (vi) Show the valence shell electronic configuration for the following:
 - a. Alkali metals s^1
 - b. Nobel gases s^2, p^6
 - c. Halogens s^2, p^5
- (vii) Determine the location of Cr^{4+} in periodic table. $6 = 2, 2 = 2$
- (viii) Briefly explain the following terms: (Any two)
 - a. Melting
 - b. Freezing
 - c. Boiling
- (ix) Find the number of electrons in valence shell in the following:
 - a. Neon
 - b. Helium
 - c. Hydrogen
- (x) Calculate the molarity of a solution that contains 40g urea (NH_2CONH_2) dissolved in 500cm³ of solution
 $14 + 2 \times 1 + 12 + 16 + 14 + 16 + 1 \times 2$
- (xi) The atomic number of an element is 20 and mass number is 40. How many Protons, Neutrons and Electrons does an atom of this element have?
- (xii) Enlist the possible uses of an electrolytic cell.
- (xiii) Find oxidation state of Nitrogen in any two of given compounds:
 - a. N_2O_5
 - b. HNO_3
 - c. NO_2
- (xiv) Magnesium reacts with steam, but calcium reacts with cold water. Show this diversity in reactivity by chemical equations.
- (xv) Why fluorine has the highest oxidizing power? Explain briefly.

SECTION - C (Marks 20)

(2 × 10 = 20)

Note: Attempt any TWO questions. All questions carry equal marks.

- Q. 3 a. What is meant by evaporation? The water level in an aquarium decreases slowly, even though the tank doesn't leak. Interpret
- b. Compare the effects of Al_2O_3 and Fe_2O_3 formation on their parent metals. (Al and Fe)

- Q. 4 a. KCl and MgF_2 are two ionic compounds. Discuss the bond formation in each of these. Also draw electron dot and cross models.
- b. Briefly discuss the following types of solutions. also give one example of each:

	Solute	Solvent
I.	Solids	Liquids
II.	Solids	Solids

- Q. 5 a. What are unsaturated, saturated and supersaturated solutions? Also give their comparison.
- b. Describe the formation of anions in Bromine and Phosphorous ions.

SUPPLEMENTARY TABLE

— 15A-1 2308 HA — 09280208

Atomic No	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20
Symbol	H	He	Li	Be	B	C	N	O	F	Ne	Na	Mg	Al	Si	P	S	Cl	Ar	K	Ca
Mass No	1	4	7	9	11	12	14	16	19	20	23	24	27	28	31	32	35.5	40	39	40



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