Balancing Equations Worksheet II

Count the number of atoms, then use coefficients to balance each equation.

AlBr₃ + ___ K → ___ KBr + ___ Al

2) ____ FeO + ___ PdF₂ → ___ FeF₂ + ___ PdO

3) ____ P₄ + ____ Br₂ → ____ PBr₃

4) ____ LiCl + ___ Br₂ → ___ LiBr + ___ Cl₂

5) ____ PbBr₂ + ___ HCl → ___ HBr + ___ PbCl₂

6) ____ CoBr₃ + ____ CaSO₄ → ____ CaBr₂ + ____ Co₂(SO₄)₃

7) ____ Na₃P + ___ CaF₂ → ___ NaF + ___ Ca₃P₂

8) ____ Mn + ___ HI → ___ H₂ + ___ MnI₃

9) ____ Li₃PO₄ + ____ NaBr → ____ Na₃PO₄ + ____ LiBr

10) ____ CaF₂ + ____ Li₂SO₄ → ____ CaSO₄ + ____ LiF

11) ____ HBr + ___ Mg(OH)₂ → ___ MgBr₂ + ___ H₂O

12) ____ LiNO₃ + ____ CaBr₂ → ____ Ca(NO₃)₂ + ____ LiBr

13) ____ AgNO₃ + ___ Li → ___ LiNO₃ + ___ Ag

14) ____ Si(OH)₄ + ___ NaBr → ___ SiBr₄ + ___ NaOH

15) ____ NaCN + ___ CuCO₃ → ___ Na₂CO₃ + ___ Cu(CN)₂

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exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students Practice of Chemistry Donald J. 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