

DESCRIBING CHEMICAL REACTIONS

Section Review - Objectives

- Explain how to write a word equation
- Describe how to write a skeleton equation
- List the steps for writing a complete chemical equation

Vocabulary

- chemical equation
- skeleton equation
- catalyst
- coefficients
- balanced equation

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

A chemical reaction can be concisely represented by a chemical ____1____. The substances that undergo a chemical change are the ____2____. The new substances formed in a chemical reaction are the ____3____. In accordance with the law of conservation of ____4____, a chemical equation must be balanced. When balancing an equation, you place ____5____ in front of reactants and products so that the same number of atoms of each ____6____ are on each side of the equation. An equation must never be balanced by changing the ____7____ in the chemical formula of a substance.

Special symbols are used to show the physical state of a substance in a reaction. The symbol for a liquid is ____8____; for a solid, ____9____; and for a gas, ____10____. A substance dissolved in water is designated ____11____. If a ____12____ is used to increase the rate of a chemical reaction, its formula is written above the arrow.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____
11. _____
12. _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 13. In an equation, a substance is shown to be in the gaseous state by placing an upward-pointing arrow after its formula.
- _____ 14. The symbol Δ placed over the arrow in an equation means that heat is supplied to the reaction.
- _____ 15. Atoms are destroyed in a chemical reaction.
- _____ 16. A skeleton equation is not a balanced equation.

Part D Questions and Problems

24. Write a balanced equation for each of these chemical reactions. I
- a. Aluminum reacts with aqueous hydrochloric acid to form hydrogen gas and aqueous aluminum chloride.
 - b. Acetylene gas (C_2H_2) burns in a welding torch with oxygen to form carbon dioxide gas and water vapor.

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John C. Kotz, Paul Treichel, Gabriela C. Weaver



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Prentice Hall Science Explorer: Chemical Interactions Pearson/Prentice Hall, Michael J. Padilla, 2004-10 *Focus on Physical Science California Edition* Michael J. Padilla, 2007 **Introductory Chemistry: Concepts and Critical Thinking** Charles H. Corwin, 2013-08-27 Were you looking for the book with access to MasteringChemistry This product is the book alone and does NOT come with access to MasteringChemistry Buy the book and access card package to save money on this resource With an expanded focus on critical thinking and problem solving the new Seventh Edition of Introductory Chemistry Concepts and Critical Thinking prepares students for success in Introductory Chemistry courses Unlike other introductory chemistry texts all materials the textbook student solutions manual laboratory manual instructor s manual and test item file are written by the author and tightly integrated to work together most effectively Math and problem solving are covered early in the text Corwin builds student confidence and ability through innovative pedagogy and technology formulated to meet the needs of today s learners By presenting chemistry in a clear and interesting way students to leave their first chemistry course with a positive impression a set of new skills and the desire to learn more **Chemistry & Chemical Reactivity** John C. Kotz, Paul Treichel, Gabriela C. Weaver, 2006 Chapter Goals and Chapter Goals Revisited are two new features in this revision Each chapter starts with a list of goals that allows students to see what is ahead The chapter concludes with a repetition of that list with summary information added General ChemistryNow is correlated to this list New to this edition are dozens of Active Figures to help students visualize chemistry in action These animated versions of text art help students master key concepts from the book Active Figures can be used as demonstrations in the classroom and each figure is paired with a guided exploration and exercise to ensure students understand the concept being illustrated In text worked Examples follow a four part structure Problem statement Strategy for approaching the problem fully worked Solution and where appropriate a Comment on the problem and solution Through this approach students learn how to approach a problem rather than merely learning to memorize problem types and memorized solution approaches Exercises appear throughout the text so students can check their comprehension of the material Answers are in an appendix Problem Solving Tips provide readers tips for determining how to approach and solve problems Chemical Perspectives are essays that bring relevance and perspective to a study of chemistry In order to put chemistry in its historical context Historical Perspective essays describe the people who were key to developing the concepts of the chapter A Closer Look essays describe ideas that form the background to material under discussion or provide another dimension of the subject Publisher **Chemical Reactions Science Learning Guide** NewPath Learning, 2014-03-01 The Chemical Reactions Student Learning Guide includes self directed readings easy to follow illustrated explanations guiding questions inquiry based activities a lab investigation key vocabulary review and assessment review questions along with a post test It covers the following standards aligned concepts Changes of Matter Chemical Reactions Formulas Balancing Equations Types of

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1. Understanding the eBook 1assessment Pearson Key Describing Chemical Reactions
 - The Rise of Digital Reading 1assessment Pearson Key Describing Chemical Reactions
 - Advantages of eBooks Over Traditional Books
2. Identifying 1assessment Pearson Key Describing Chemical Reactions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 1assessment Pearson Key Describing Chemical Reactions
 - User-Friendly Interface
4. Exploring eBook Recommendations from 1assessment Pearson Key Describing Chemical Reactions
 - Personalized Recommendations
 - 1assessment Pearson Key Describing Chemical Reactions User Reviews and Ratings

- 1assessment Pearson Key Describing Chemical Reactions and Bestseller Lists
- 5. Accessing 1assessment Pearson Key Describing Chemical Reactions Free and Paid eBooks
 - 1assessment Pearson Key Describing Chemical Reactions Public Domain eBooks
 - 1assessment Pearson Key Describing Chemical Reactions eBook Subscription Services
 - 1assessment Pearson Key Describing Chemical Reactions Budget-Friendly Options
- 6. Navigating 1assessment Pearson Key Describing Chemical Reactions eBook Formats
 - ePub, PDF, MOBI, and More
 - 1assessment Pearson Key Describing Chemical Reactions Compatibility with Devices
 - 1assessment Pearson Key Describing Chemical Reactions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 1assessment Pearson Key Describing Chemical Reactions
 - Highlighting and Note-Taking 1assessment Pearson Key Describing Chemical Reactions
 - Interactive Elements 1assessment Pearson Key Describing Chemical Reactions
- 8. Staying Engaged with 1assessment Pearson Key Describing Chemical Reactions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 1assessment Pearson Key Describing Chemical Reactions
- 9. Balancing eBooks and Physical Books 1assessment Pearson Key Describing Chemical Reactions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 1assessment Pearson Key Describing Chemical Reactions
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 1assessment Pearson Key Describing Chemical Reactions
 - Setting Reading Goals 1assessment Pearson Key Describing Chemical Reactions
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 1assessment Pearson Key Describing Chemical Reactions
 - Fact-Checking eBook Content of 1assessment Pearson Key Describing Chemical Reactions
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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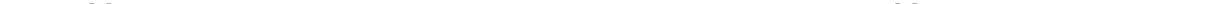
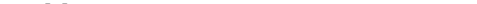
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