

Name: _____ ANSWER KEY _____ Class: _____ Date: _____

Unique Properties of Water!

Word Bank:

Adhesion	Evaporation	Polar	Surface tension
Cohesion	Freezing	Positive	Universal solvent
Condensation	Melting	Sublimation	
Dissolve	Negative		

1. The electrons are not shared equally between the hydrogen and oxygen atoms of water creating a Polar molecule.
2. The polarity of water allows it to dissolve most substances. Because of this it is referred to as the universal solvent.
3. Water molecules stick to other water molecules. This property is called cohesion.
4. Hydrogen bonds form between adjacent water molecules because the positive charged hydrogen end of one water molecule attracts the negative charged oxygen end of another water molecule.
5. Water molecules stick to other materials due to its polar nature. This property is called adhesion.
6. Hydrogen bonds hold water molecules closely together which causes water to have high surface tension. This is why water tends to clump together to form drops rather than spread out into a thin film.
7. Condensation is when water changes from a gas to a liquid.
8. Sublimation is when water changes from a solid directly to a gas.
9. Freezing is when water changes from a liquid to a solid.
10. Melting is when water changes from a solid to a liquid.
11. Evaporation is when water changes from a liquid to a gas.
12. Why does ice float?
Water expands as it freezes, so it is LESS DENSE AS A SOLID.
13. What property refers to water molecules resembling magnets? How are these alike?
Polar bonds create positive and negative ends of the molecule. The positive ends are attracted to other negative ends, creating hydrogen bonds between molecules.

2 Properties Of Water Answers

RJ Shavelson



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The Properties of Water and their Role in Colloidal and Biological Systems Carel Jan van Oss, 2008-09-16 This book treats the different current as well as unusual and hitherto often unstudied physico chemical and surface thermodynamic properties of water that govern all polar interactions occurring in it These properties include the hyper hydrophobicity of the water air interface the cluster formation of water molecules in the liquid state and the concomitant variability of the ratio of the electron acceptivity to electron donicity of liquid water as a function of temperature T The increase of that ratio with T is the cause of the increase in hydration repulsion hydration pressure between polar surfaces upon heating when they are immersed in water The book also treats the surface properties of apolar and polar molecules polymers particles and cells as well as their mutual interaction energies when immersed in water under the influence of the three prevailing non covalent forces i e Lewis acid base AB Lifshitz van der Waals LW and electrical double layer EL interactions The polar AB interactions be they attractive or repulsive typically represent up to 90% of the total interaction energies occurring in water Thus the addition of AB energies to the LW EL energies of the classical DLVO theory of energy vs distance analysis makes this powerful tool the Extended DLVO theory applicable to the quantitative study of the stability of particle suspensions in water The influence of AB forces on the interfacial tension between water and other condensed phase materials is stressed and serves inter alia to explain measure and calculate the driving force of the hydrophobic attraction between such materials the hydrophobic effect when immersed in water These phenomena which are typical for liquid water influence all polar interactions that take place in it All of these are treated from the viewpoint of the properties of liquid water itself including the properties of advancing freezing fronts and the surface properties of ice at 0o C Explains and allows the quantitative measurement of hydrophobic attraction and hydrophilic repulsion in water Measures the degree of cluster formation of water molecules Discusses the influence of temperature on the cluster size of water molecules Treats the multitudinous effects of the hyper hydrophobicity of the water air interface Chemistry in the Community American Chemical Society, 2006-01-31 This laboratory based text centres itself around decision making activities where students apply their chemistry knowledge to realistic situations This fifth edition includes more photographs new drawings and new design *Essential Cell Biology* Mr. Rohit Manglik, 2024-07-25 EduGorilla Publication is a trusted name in the education sector committed to empowering learners with high quality study materials and resources Specializing in competitive exams and academic support EduGorilla

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