

Use the activity series of elements to predict which reactions take place.

Metal	Metal Ions			
	Copper (Cu)	High (H)	Low (L)	Ag (Ag)
Copper (Cu)	NO	NO More active	YES More active	NO
Magnesium (Mg)	YES	NO	NO	NO
Zinc (Zn)	YES	NO More active	NO	NO
Silver (Ag)	NO More active	NO More active	YES More active	NO

Questions

1. Write your cell as a spontaneous reaction with a balanced equation using this activity series as a reference.



2. Which metal reacts with the most solutions? **Magnesium**

3. Which metal reacts with the least solutions? **Ag**

4. List three metals in order of greatest to least reactivity from the metals that reacted with the most solutions and substituting in order of the least reactivity metal that reacted with the least solutions. Rank is higher if found to be an active metal reacting quickly in acidic metals.



5. Write the net ionic equation for a reaction of a metal with an aqueous solution of a metal ion. Write the net ionic equation for the reaction of a metal with an aqueous solution of a metal ion.



6. Write a balanced net ionic equation for the reaction of a metal with an aqueous solution of a metal ion.



7. Write a balanced net ionic equation for the reaction of a metal with an aqueous solution of a metal ion.

An Activity Series Lab Answers

C. Alton Hassell, Paula Marshall



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Table of Contents An Activity Series Lab Answers

1. Understanding the eBook An Activity Series Lab Answers
 - The Rise of Digital Reading An Activity Series Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying An Activity Series Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an An Activity Series Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from An Activity Series Lab Answers
 - Personalized Recommendations
 - An Activity Series Lab Answers User Reviews and Ratings
 - An Activity Series Lab Answers and Bestseller Lists
5. Accessing An Activity Series Lab Answers Free and Paid eBooks
 - An Activity Series Lab Answers Public Domain eBooks
 - An Activity Series Lab Answers eBook Subscription Services
 - An Activity Series Lab Answers Budget-Friendly Options

6. Navigating An Activity Series Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - An Activity Series Lab Answers Compatibility with Devices
 - An Activity Series Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of An Activity Series Lab Answers
 - Highlighting and Note-Taking An Activity Series Lab Answers
 - Interactive Elements An Activity Series Lab Answers
8. Staying Engaged with An Activity Series Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers An Activity Series Lab Answers
9. Balancing eBooks and Physical Books An Activity Series Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection An Activity Series Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine An Activity Series Lab Answers
 - Setting Reading Goals An Activity Series Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of An Activity Series Lab Answers
 - Fact-Checking eBook Content of An Activity Series Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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