

**ホンダ**  
**CBR250FOUR**  
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**サービスマニュアル**



G~R

# 1987 1991 Honda Cbr250 Motorcycle Service Manual Japanese

**IM Harris**



## **1987 1991 Honda Cbr250 Motorcycle Service Manual Japanese:**

**Honda Service Manual 87-89 CBR, CBR600F.** ,1988      **Honda Service Manual[published 1985 Through 1991].**  
,1985      **Honda Service Manual, 84-87** Honda Giken Kōgyō Kabushiki Kaisha,1986      **Honda Service Manual, 85-87**  
,1986      **Honda Service Manual [published 1991 Through 1992].** ,1991      *Honda Service Manual [published 1987*  
*Through 1988].* ,1987      **Honda Service Manual, 86-91 CR250R.** ,1990      *Honda Service Manual 87-90, 92-95 VT1100C*  
*Shadow* ,1994      **Honda Service Manual 91-94 CBR600F2** Honda Giken Kōgyō Kabushiki Kaisha,1993      **Honda**  
**CBR125R, CBR250R & CRF250L/M, '11-'14** Haynes Publishing,2020-02-25 Each Haynes manual provides specific and  
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1991 1994 U S CBR600F3 1995 1998 U S CBR600SJR Smokin Joe s 1996 U S CBRSE Smokin Joe s 1998 U S CBR600F U K

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 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$   $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$   $\Delta h = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$   $\Delta h = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + 1/2 \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 285 \text{ kJ/mole}$

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 $\text{C}_2\text{H}_4(\text{g}) + \text{NO}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + \text{O}_2(\text{g}) + \text{N}_2\text{O}_5(\text{g})$   
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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{SO}_4$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta h = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta h_{\text{products}} > \Delta h_{\text{reactants}}$

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