

Rule 3: The greater the charge on the ions, the greater is the polarization produced and thereby the greater is the covalent character of the bond.

Explanation: A high charge on the cation or the anion or both will obviously increase the attraction between the cation and the electron cloud of the anion. This will clearly increase the deformation of the anion.

For example, consider NaCl, MgCl₂, and AlCl₃. The higher charge on Al³⁺ than on the other two (Na, Mg) renders Al³⁺ more capable than the other two in polarizing the anion (chloride ion). Consequently, AlCl₃ has the highest covalent character and NaCl has the least.

Rule 4: Among cations of the same charge and comparable sizes, that having the pseudo-inert gas configuration ns² np⁶ nd¹⁰ has a higher polarizing power than that with the noble gas configuration ns² np⁶.

Explanation: The shielding of the nuclear charge in an inert gas configuration is very effective while that in a pseudo-inert gas configuration is very poor. This is due to the fact that the ten d electrons in a pseudo inert gas type ion will shield its nuclear charge less effectively than s and p electrons.

For example, consider Cu⁺, Ag⁺ and Au⁺. They have ns² np⁶ nd¹⁰ configurations which are very poor at shielding the nuclear charge; so they are highly polarizing and induce high covalent character into bonds with anions. Their chlorides (CuCl, AgCl and AuCl) possess considerable covalent character.

The ns² np⁶ configurations of Na⁺, K⁺ and Pb²⁺ are very effective at shielding the nuclear charge so that they are poorly polarizing and induce very little covalent character into bonds with anions. NaCl, KCl, and PbCl₂ are ionic compounds with very little covalent character.

Shielding (Screening Effect)

A negatively charged electron can shield the shielding of an electron with respect to the attraction of the nucleus and consequently the net of the forces is called a shielding force or screening force or a shielding effect or a screening power.

The screening effect is due to the shielding of the nucleus by electrons with the same effect as the shielding of the nucleus by the shielding of the nucleus by the nucleus. The screening force due to the shielding of the nucleus is called a shielding force or a screening power or a screening effect or a screening power.

Example:



Consider the attractive forces between the nucleus and the electrons. The attractive force of the nucleus on the electrons is called a shielding force or a screening force or a screening power or a screening effect or a screening power.

8 2 Ws Answers The Nature Of Covalent

Nima Rezaei



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