

## Chapter 4 Atomic Structure

**Section 4.1 Studying Atoms****(pages 100-105)***This section discusses the development of atomic models.***Reading Strategy (page 100)**

**Summarizing** As you read, complete the table about atomic models. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Atomic Models		
Scientist	Evidence	Model
	Ratio of masses in compounds	
	Deflected beam	
Rutherford		Positive, dense nucleus

**Ancient Greek Models of Atoms (page 100)**

- Democritus named the smallest particles of matter \_\_\_\_\_ because they could not be divided.
- List the four elements that Aristotle included in his model of matter.
  - \_\_\_\_\_
  - \_\_\_\_\_
  - \_\_\_\_\_
  - \_\_\_\_\_

**Dalton's Atomic Theory (page 101)**

- Is the following sentence true or false? John Dalton gathered evidence for the existence of atoms by measuring the masses of elements that reacted to form compounds. \_\_\_\_\_
- What theory did Dalton propose to explain why the elements in a compound always join in the same way? \_\_\_\_\_  
\_\_\_\_\_
- Circle the letters of the sentences that represent the main points of Dalton's theory of atoms.
  - All elements are composed of atoms.
  - In a particular compound, atoms of different elements always combine the same way.
  - All atoms have the same mass.
  - Compounds contain atoms of more than one element.

## 4 1 Studying Atoms Answer Key

**Edward Maslowsky, Jr.**



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