

Question number	Answer	Marks	Guidance
1 a i	1.6734×10^{-26} (g) 1.6734×10^{-27} kg	1	Only. Not 1.67×10^{-24} (g)
1 a ii	B	1	
1 b i	$\frac{10x + 11y}{x + y} = 10.8$ OR ratio 10:11 = 1:4 OR 20:80 etc abundance of ^{10}B is 20(%) OR $\frac{10x}{100} + \frac{11(100 - x)}{100} = 10.8$ $10x + 1100 - 11x = 1080$ $\therefore x = 1100 - 1080$ $= 20\%$	1 1	Allow idea that there are 5×0.2 divisions between 10 and 11. Correct answer scores M1 and M2.
1 b ii	Same number of electrons (in outer shell or orbital) Same electronic configuration / arrangement	1	Ignore electrons determine chemical properties. Ignore protons unless wrong.
1 c	Range between 3500 and 10 000 kJ mol ⁻¹	1	Ignore electrons determine chemical properties. Ignore protons unless wrong.
1 d	$\text{B}^+(g) \rightarrow \text{B}^{2+}(g) + e^{-1}$ $\text{B}^+(g) - e^{-1} \rightarrow \text{B}^{2+}(g)$ $\text{B}^+(g) + e^{-1} \rightarrow \text{B}^{2+}(g) + 2e^{-1}$	1	Ignore state symbol on electron even if wrong.
1 e	Electron being removed from a positive ion (therefore need more energy) / electron being removed is closer to the nucleus.	1	Must imply removal of an electron. Allow electron removed from a + particle/ species or from a 2+ ion. Not electron removed from a higher / lower energy level / shell. Not electron removed from a higher energy sub-level / orbital. Ignore electron removed from a lower energy sub-level / orbital. Ignore 'more protons than electrons'. Not 'greater nuclear charge'. Ignore 'greater effective nuclear charge'. Ignore shielding.
2	$\frac{(82.0 \times 2) + (83.0 \times 2) + (84.0 \times 10) + (86.0 \times 3)}{17.0}$ $\frac{1428}{17.0}$ $= 84.0$ The <i>A</i> in the Periodic table takes account of the other	2 1 1	M1 for the top line. M2 is for division by 17. Not 84 Award independently.

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Lingsheng Yao

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