

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PerV-nRT", and the universal gas constant $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 kPa / 1 atm$ to get
 $R = 8.31 kPa \cdot L / (K \cdot mole)$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 atm)(12 L)}{4 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol}}$$

$$T = 204.63 K$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87 °C, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 atm \cdot 31 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 K}$$

$$n = 1.2586 mol$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 400 K}{60 L}$$

$$P = 1.642 atm$$

- 4) If I have 1.7 moles of gas at a pressure of 0.95 atm and at a temperature of 36 °C, what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{1.7 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 309 K}{0.95 atm}$$

$$V = 47.169 L$$

- 5) If I have 17 moles of gas at a temperature of 87 °C, and a volume of 88.99 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 K}{88.99 L}$$

$$P = 5.594 atm$$

- 6) If I have an unknown quantity of gas at a pressure of 0.9 atm, a volume of 29 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.9 atm \cdot 29 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 300 K}$$

$$n = 0.5075 mol$$

Ap Chemistry Gas Laws Worksheet

Philip Hicks

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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