

Ideal Gas Law Worksheet PV = nRT

Use the ideal gas law, "PerV-nRT", and the universal gas constant $R = 0.0821 \frac{L.*atm}{K^*mol}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by 101.3kPa / latm to get

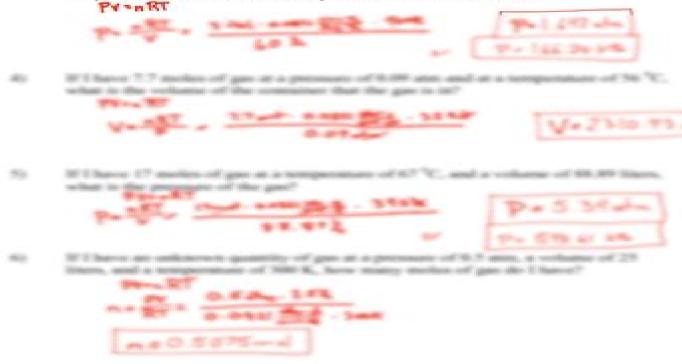
R=8.31 kPa*L / (K*mole)

1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?



2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87 °C, how many moles of gas do I have?

3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?



Ap Chemistry Gas Laws Worksheet

Philip Hicks

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The Gas Laws Irene Cesa, 2003-01-01.

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