

AP Chemistry - Chemical Kinetics

Name: _____ Date: _____ Period: _____

- After five half-life periods for a first order reaction, what fraction of reactant remains?
- If you plot $1/[\text{reactant}]$ versus time and observe a straight line, what is the order of the reaction? If $\ln[\text{reactant}]$ is plotted versus time, and a straight line of negative slope is observed, what is the order of the reaction?
- Indicate whether each of the following statements is true or false. Change the wording of each false statement to make it true.
 - In stepwise reactions the rate-determining step is the slow one.
 - It is possible to change the rate constant for a reaction by changing the temperature.
 - As a first order reaction proceeds at a constant temperature, the rate remains constant.
 - The rate constant for a reaction is independent of reactant concentrations.
- For each reaction below give the relative rates of formation of products and disappearance of reactants using coefficients and rate change symbols.
 - $2\text{O}_{3(g)} \rightarrow 3\text{O}_{2(g)}$
 - $2\text{HOF}_{(g)} \rightarrow 2\text{HF}_{(g)} + \text{O}_{2(g)}$
 - $\text{N}_{2(g)} + 3\text{H}_{2(g)} \rightarrow 2\text{NH}_{3(g)}$
- The reaction $\text{CO}_{(g)} + \text{NO}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{NO}_{(g)}$ is second order in NO_2 and zero order in CO at temperatures less than 500K.
 - Write the rate law for the reaction.
 - How will the reaction rate change if the NO_2 concentration is halved?
- For a reaction of the type $\text{A} + \text{B} \rightarrow \text{C}$, it is experimentally observed that doubling the concentration of B causes the reaction rate to be increased fourfold, but doubling the concentration of A has no effect on the rate. Determine the initial rate equation.
- A study of the reaction $2\text{A} + \text{B} \rightarrow \text{C} + \text{D}$ gave the following results:

Expt	Init Con (M)		Init Rate (M/s)
	[A] ₀	[B] ₀	
1.	0.10	0.05	6.0×10^{-3}
2.	0.20	0.05	1.2×10^{-2}
3.	0.30	0.05	1.8×10^{-2}
4.	0.20	0.15	1.1×10^{-1}

- What is the rate law for this reaction?
 - What is the overall order of the reaction?
 - Could this reaction occur in a single step?
- The decomposition of N_2O_5 (to give NO_2 and O_2) follows the rate law "rate = $k[\text{N}_2\text{O}_5]^n$ " where k is 5.0×10^{-4} per second at a particular temperature.
 - What is the half-life of this reaction?
 - How long does it take for the N_2O_5 concentration to drop to one tenth of its original value?

Ap Chemistry Chemical Kinetics Worksheet Answers

John T. Moore, Richard H. Langley



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