

AP CHEMISTRY CHAPTER 6 PRACTICE TEST

- For which chemical equation will the heat of reaction ΔH° , correspond to the molar heat of formation of liquid water?
 - $\text{H}_{2(g)} + 1/2\text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(l)}$
 - $\text{H}_{2(g)} + \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(l)}$
 - $2\text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2\text{H}_2\text{O}_{(l)}$
 - $\text{H}^+_{(aq)} + \text{OH}^-_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)}$
- For the process $\text{H}_{2(g)} + 1/2\text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(l)}$, $\Delta H = -247 \text{ kJ}$. The numeric value of -242 kJ represents heat energy that is
 - absorbed by the reaction
 - required to break the H-H and O=O bond
 - stored as potential energy
 - released by the reaction
- An 11.0 g sample of propane is placed in a bomb calorimeter with excess oxygen and is ignited. The initial temperature of the calorimeter is 25.00°C and its total heat capacity is $97.1 \text{ kJ}/^\circ\text{C}$. The reaction raises the temperature of the calorimeter is 30.75°C . The approximate molar heat of combustion for propane is
 - -50.6 kJ/mol
 - -558 kJ/mol
 - $-2.24 \times 10^3 \text{ kJ/mol}$
 - $-1.20 \times 10^4 \text{ kJ/mol}$
- What is the change in enthalpy at 25°C and 1 atm for the production of 3.00 mol $\text{SnO}_{(s)}$?

$$\text{Sn}_{(s)} + \text{SnO}_{2(s)} \rightarrow 2\text{SnO}_{(s)} \quad \Delta H^\circ = 16.2 \text{ kJ}$$
 - 24.3 kJ
 - 16.2 kJ
 - 5.40 kJ
 - -16.2 kJ
 - -24.3 kJ
- How much heat is liberated at constant pressure if 0.896 g of calcium carbonate reacts with 92.3 mL of 0.584 M hydrochloric acid?

$$\text{CaCO}_{3(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{CaCl}_{2(aq)} + \text{H}_2\text{O}_{(l)} + \text{CO}_{2(g)} \quad \Delta H^\circ = -15.2 \text{ kJ}$$
 - -0.410 kJ
 - -8.88 kJ
 - -13.6 kJ
 - -0.546 kJ
 - -0.136 kJ

Ap Chemistry Chapter 6 Practice Test

BM King



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