

AP Chapter 2 Study Questions**True/False**

Indicate whether the statement is true or false.

1. The least electronegative halogen is astatine.
2. The possible oxidation numbers for iron are +1 and +2.
3. The formula for chromium (II) iodide is CrI_2 .
4. H_2SeO_4 is called selenic acid.
5. The correct name for Na_3N is sodium azide.

Multiple Choice

Identify the choice that best completes the statement or answers the question.

6. A certain mass of carbon reacts with 13.6 g of oxygen to form carbon monoxide. _____ grams of oxygen would react with that same mass of carbon to form carbon dioxide, according to the law of multiple proportions?
a. 25.6 b. 6.8 c. 13.6 d. 136 e. 27.2
7. Methane and ethane are both made up of carbon and hydrogen. In methane, there are 12.0 g of carbon for every 4.00 g of hydrogen, a ratio of 3:1 by mass. In ethane, there are 24.0 g of carbon for every 6.00 g of hydrogen, a ratio of 4:1 by mass. This is an illustration of the law of _____.
a. constant composition b. multiple proportions
c. conservation of matter d. conservation of mass e. octaves
8. Which statement below correctly describes the responses of alpha, beta, and gamma radiation to an electric field?
a. Both beta and gamma are deflected in the same direction, while alpha shows no response. b. Both alpha and gamma are deflected in the same direction, while beta shows no response. c. Both alpha and beta are deflected in the same direction, while gamma shows no response. d. Alpha and beta are deflected in opposite directions, while gamma shows no response. e. Only alpha is deflected, while beta and gamma show no response.
9. _____ and _____ reside in the atomic nucleus.
a. Protons, electrons b. Electrons, neutrons
c. Protons, neutrons d. none of the above
e. Neutrons, only neutrons
10. 200 pm is the same as _____ Å.
a. 2000 b. 20 c. 200 d. 2 e. 2×10^{-12}
11. The atomic number indicates _____.
a. the number of neutrons in a nucleus b. the total number of neutrons and protons in a nucleus
c. the number of protons or electrons in a neutral atom d. the number of atoms in 1 g of an element
e. the number of different isotopes of an element
12. Which pair of atoms constitutes a pair of isotopes of the same element?
a. $\begin{smallmatrix} 14\text{X} \\ 6 \end{smallmatrix}$ $\begin{smallmatrix} 14\text{X} \\ 7 \end{smallmatrix}$ b. $\begin{smallmatrix} 14\text{X} \\ 6 \end{smallmatrix}$ $\begin{smallmatrix} 12\text{X} \\ 6 \end{smallmatrix}$ c. $\begin{smallmatrix} 17\text{X} \\ 9 \end{smallmatrix}$ $\begin{smallmatrix} 17\text{X} \\ 8 \end{smallmatrix}$
d. $\begin{smallmatrix} 19\text{X} \\ 10 \end{smallmatrix}$ $\begin{smallmatrix} 19\text{X} \\ 9 \end{smallmatrix}$ e. $\begin{smallmatrix} 20\text{X} \\ 10 \end{smallmatrix}$ $\begin{smallmatrix} 21\text{X} \\ 11 \end{smallmatrix}$
13. The nucleus of an atom contains _____.
a. electrons b. protons, neutrons, and electrons
c. protons and neutrons d. protons and electrons
e. protons

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CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

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