

## AP Chemistry: Chapter 14 Practice Test

10. Given the following acids and  $K_a$  values:

$\text{HClO}_4$	$\text{HOAc}$	$\text{HCN}$	$\text{HF}$
-----------------	---------------	--------------	-------------

$1 \times 10^7$	$1.76 \times 10^{-5}$	$4.93 \times 10^{-10}$	$3.53 \times 10^{-4}$
-----------------	-----------------------	------------------------	-----------------------

which shows the conjugate bases listed by increasing strength?

- |  |  |
|--|--|
| [A] $\text{ClO}_4^-$ , $\text{F}^-$ , $\text{OAc}^-$ , $\text{CN}^-$ | [B] $\text{CN}^-$ , $\text{F}^-$ , $\text{OAc}^-$ , $\text{ClO}_4^-$ |
| [C] $\text{CN}^-$ , $\text{OAc}^-$ , $\text{F}^-$ , $\text{ClO}_4^-$ | [D] $\text{CN}^-$ , $\text{ClO}_4^-$ , $\text{F}^-$ , $\text{OAc}^-$ |
| [E] $\text{ClO}_4^-$ , $\text{OAc}^-$ , $\text{CN}^-$ , $\text{F}^-$ |  |
15. Given that the  $K_a$  for  $\text{HOCl}$  is  $3.5 \times 10^{-8}$ , calculate the  $K$  value for the reaction of  $\text{HOCl}$  with  $\text{OH}^-$ .
- |                           |                       |                          |
|---------------------------|-----------------------|--------------------------|
| [A] $3.5 \times 10^{-22}$ | [B] $3.5 \times 10^6$ | [C] $2.9 \times 10^{-7}$ |
| [D] $3.5 \times 10^{-8}$  | [E] none of these     |                          |
32. Acetic acid ( $\text{HC}_2\text{H}_3\text{O}_2$ ) is a weak acid ( $K_a = 1.8 \times 10^{-5}$ ). Calculate the pH of a 17.6 M  $\text{HC}_2\text{H}_3\text{O}_2$  solution.
- |         |         |         |          |         |
|---------|---------|---------|----------|---------|
| [A] 4.3 | [B] 6.4 | [C] 1.7 | [D] 0.97 | [E] 7.4 |
|---------|---------|---------|----------|---------|
52. What concentration of acetic acid ( $K_a = 1.80 \times 10^{-5}$ ) has the same pH as that of  $5.00 \times 10^{-2}$  M  $\text{HCl}$ ?
- |                             |                             |            |
|-----------------------------|-----------------------------|------------|
| [A] $5.00 \times 10^{-2}$ M | [B] $3.60 \times 10^{-3}$ M | [C] 1.39 M |
| [D] 5.00 M                  | [E] none of these           |            |

Consider the reaction  $\text{HOCl} + \text{F}^- \rightleftharpoons \text{HF} + \text{OCl}^-$

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**Peter J. Mikulecky,Michelle Rose  
Gilman,Kate Brutlag**



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### **Table of Contents Ap Chemistry Chapter 14 Practice Test**

1. Understanding the eBook Ap Chemistry Chapter 14 Practice Test
  - The Rise of Digital Reading Ap Chemistry Chapter 14 Practice Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Chapter 14 Practice Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ap Chemistry Chapter 14 Practice Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Chapter 14 Practice Test
  - Personalized Recommendations
  - Ap Chemistry Chapter 14 Practice Test User Reviews and Ratings
  - Ap Chemistry Chapter 14 Practice Test and Bestseller Lists

5. Accessing Ap Chemistry Chapter 14 Practice Test Free and Paid eBooks
  - Ap Chemistry Chapter 14 Practice Test Public Domain eBooks
  - Ap Chemistry Chapter 14 Practice Test eBook Subscription Services
  - Ap Chemistry Chapter 14 Practice Test Budget-Friendly Options
6. Navigating Ap Chemistry Chapter 14 Practice Test eBook Formats
  - ePub, PDF, MOBI, and More
  - Ap Chemistry Chapter 14 Practice Test Compatibility with Devices
  - Ap Chemistry Chapter 14 Practice Test Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ap Chemistry Chapter 14 Practice Test
  - Highlighting and Note-Taking Ap Chemistry Chapter 14 Practice Test
  - Interactive Elements Ap Chemistry Chapter 14 Practice Test
8. Staying Engaged with Ap Chemistry Chapter 14 Practice Test
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ap Chemistry Chapter 14 Practice Test
9. Balancing eBooks and Physical Books Ap Chemistry Chapter 14 Practice Test
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ap Chemistry Chapter 14 Practice Test
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ap Chemistry Chapter 14 Practice Test
  - Setting Reading Goals Ap Chemistry Chapter 14 Practice Test
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ap Chemistry Chapter 14 Practice Test
  - Fact-Checking eBook Content of Ap Chemistry Chapter 14 Practice Test
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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