

Thermodynamics Quiz #1 Answer Key

2004

$$(a) \Delta G^\circ = [\Delta_f G^\circ] + [0] - [-182.47 + 0] = 182.7 \text{ kJ/mol}$$

$$K_c = e^{\frac{\Delta G^\circ}{RT}} = e^{\frac{182.7 \times 10^3}{8.314 \times 298.15}} = 1.80 \times 10^{-4}$$

$$(b) \Delta H^\circ = [\Delta_f H^\circ] + [0] - [-285.83 + 0] = 27.44 \text{ kJ/mol}$$

$$\Delta S^\circ = \frac{\Delta_f H^\circ - \Delta_f G^\circ}{T} = \frac{27440 - 182470}{298.15} = -29.27 \frac{\text{J}}{\text{mol}\cdot\text{K}}$$

$$\Delta G^\circ = 0 \text{ at eq. } K_c = \left(\frac{P_{\text{H}_2\text{O}}}{P_{\text{H}_2}} \right)^2 = 1 \text{ at equilibrium}$$

(c)

$$T = \Delta H^\circ / \Delta S^\circ = 27440 / 29.2 = 939 \text{ K}$$

2003 B

$$(a) \Delta_f H^\circ = \sum \Delta_f H^\circ(\text{products}) - \sum \Delta_f H^\circ(\text{reactants}) = -108.2 - [-137.2 + 2(22)] = -28.8 \text{ kJ/mol}$$

$$(b) K_c = e^{\frac{\Delta G^\circ}{RT}} = e^{\frac{-28.8 \times 10^3}{8.314 \times 298.15}} = 1.16 \times 10^5$$

$$(c) \Delta_f G^\circ = \frac{\Delta_f H^\circ - \Delta_f G^\circ}{T} = \frac{-108.2 \text{ kJ} - (-28.8 \text{ kJ})}{298 \text{ K}} = -33.5 \text{ kJ}$$

(d) Both the standard enthalpy of formation and the standard free energy of formation of acetone = 0.

$$-2 \Delta_f G^\circ_{\text{acetone}} = -\Delta_f G^\circ_{\text{ethanol}}$$

$$-33.5 \text{ kJ} = 126.8 \text{ kJ} - 187.8 \text{ kJ} - 2 \Delta_f G^\circ$$

$$\Delta_f G^\circ = 131.3 \text{ kJ/mol}$$

2004 B

$$(a) \Delta_f H^\circ = \sum \Delta_f H^\circ(\text{products}) - \sum \Delta_f H^\circ(\text{reactants}) = [4(201.5) + 4(235.85) - 2(80.3)] \text{ kJ} = -321.8 \text{ kJ}$$



$$(c) \Delta_f G^\circ(\text{products}) - \Delta_f G^\circ(\text{reactants}) = [4(57.0) + 4(57.0) + 57.0] - 236.3 - [4(5.69) + 4(130.4)] = -323.8 \text{ kJ}$$

$$(d) \Delta G^\circ = \Delta_f G^\circ - T \Delta_f G^\circ = -321.8 - (298)(-0.5230) \text{ kJ} = -375.7 \text{ kJ}$$

Ap Chemistry 20answer Key

Richard H. Langley, John Moore



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