

AP Chem Practice Test 6 – Chapters 9 & 10

1. Draw a reasonable Lewis structure for each species and sketch its shape. Include the number of valence electrons, AXE designation, and geometry & molecular shape, and bond & molecular polarity. Identify any structures that do not conform to octet rule, and describe how they violate octet rule. Give the VB hybridization scheme for the central atom, and diagram the bonds (i.e., indicate the orbitals involved and the type of overlap).



2. Explain why second period elements have never been observed as the central atom in structures with expanded octets.

3. a. Which of the following molecules would you expect to have a dipole moment, and why:



- b. Explain why O_3 has a dipole moment.

4. Draw Lewis structures, assign formal charges, and choose the most likely structure:



5. a. Draw resonance structures for NO_2 . What are the bonds really like? What would you predict for the bond order and bond lengths (be specific: numbers, please, see Table 9.4 on pg 395). How does VB theory account for resonance? [For more practice, try NO_2^{1-} and O_3]
b. Draw the resonance structures for benzene, C_6H_6 , a popular item on AP questions.

6. Use bond energies (Table 9.3 p. 392) to estimate the enthalpy of this combustion reaction:



The reported value is $-1411 \text{ kJ/mol C}_2\text{H}_4$. Evaluate the usefulness of bond energies for estimating reaction enthalpy.

7. Sketch a diagram that shows the VB model of orbital overlap for the molecule shown below. Identify σ and π bonding areas, hybridization, and bond angles. How does π bonding affect the freedom of rotation around the bond?



8. Sketch the bonding and antibonding orbitals for σ_{1s} and σ_{1s}^* MOs and describe how they differ.

9. Sketch a labeled diagram (like the one on the bottom left on pg 445) that shows the MOs for principal energy level 2. In your MO diagram, place the valence electrons for BC^{1-} . Calculate the bond order and determine its stability. Is it diamagnetic or paramagnetic? How would adding another electron affect its stability? How would adding another electron affect the bond length?

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