

- (d) In the box below, draw a complete Lewis electron-dot diagram for one valid resonance structure of $\text{SO}_2(\text{g})$.
Any one of the three following diagrams is acceptable.



One point for the correct number of electrons.

One point for a valid Lewis diagram.

1 point

3.B

SAP-4.A

Total for part (d)

2 points

- (e) Based on the diagram you drew in part (d), what is the approximate oxygen-sulfur-oxygen bond angle in SO_2 ?

SO_2 can be oxidized to form SO_3 according to the following equation.



One point for an angle that is consistent with the student's Lewis structure: 120°

1 point

1.A

SAP-4.C

- (f) Is the value of ΔS° for the reaction represented above positive or negative? Justify your answer.

One point for indicating negative with a valid explanation that mentions the smaller number of moles of gas in the products.

- Negative, because the reactants are three moles of gas but the products are only two moles of gas.

1 point

4.C

ENE-4.A

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