

AP Chemistry – Chapter 13, Chemical Equilibrium Study Guide & Ch. 16a K_{sp}

Students should be able to...

- Describe a [system in equilibrium](#)
- Find the reaction [quotient, Q](#), (K_{eq}) for a reaction equation
- Solve for Q given [initial concentrations or pressures](#)
- Using [LeChatelier's principle](#), determine the direction of shift in an equilibrium when conditions are changed (concentration, pressure, phase change)
- Setup a [K_{eq} expression for a slightly soluble product](#).
- Calculate the [ion concentrations](#) when given a [K_{sp}](#)
- Determine is a slightly soluble product will precipitate or not.

Sample Multiple Choice AP Questions (Copyright College Board)

(1989) 29. In which of the following systems would the number of moles of the substances present at equilibrium NOT be shifted by a change in the volume of the system at constant temperature?

- (A) $\text{CO(g)} + \text{NO(g)} \rightleftharpoons \text{CO}_2\text{(g)} + 1/2 \text{N}_2\text{(g)}$
- (B) $\text{N}_2\text{(g)} + 3 \text{H}_2\text{(g)} \rightleftharpoons 2 \text{NH}_3\text{(g)}$
- (C) $\text{N}_2\text{(g)} + 2 \text{O}_2\text{(g)} \rightleftharpoons 2 \text{NO}_2\text{(g)}$
- (D) $\text{N}_2\text{O}_4\text{(g)} \rightleftharpoons 2 \text{NO}_2\text{(g)}$
- (E) $\text{NO(g)} + \text{O}_2\text{(g)} \rightleftharpoons \text{NO}_2\text{(g)} + \text{O}_2\text{(g)}$

(1994) 73. $2 \text{SO}_2\text{(g)} + \text{O}_2\text{(g)} \rightleftharpoons 2 \text{SO}_3\text{(g)}$

When 0.40 mole of SO₂ and 0.60 mole of O₂ are placed in an evacuated 1.00-liter flask, the reaction represented above occurs. After the reactants and the product reach equilibrium and the initial temperature is restored, the flask is found to contain 0.30 mole of SO₃. Based on these results, the equilibrium constant, K_{eq}, for the reaction is

- (A) 20.
- (B) 10.
- (C) 6.7
- (D) 2.0
- (E) 1.2

65. The solubility of CuI is 2×10^{-8} molar. What is the solubility product constant, K_{sp}, for CuI?

- (A) 1.4×10^{-7}
- (B) 2×10^{-8}
- (C) 4×10^{-12}
- (D) 2×10^{-12}
- (E) 8×10^{-18}

66. $\text{MnS(s)} + 2 \text{H}^+ \rightleftharpoons \text{Mn}^{2+} + \text{H}_2\text{S(g)}$

At 25 °C the solubility product constant, K_{sp}, for MnS is 5×10^{-15} and the acid dissociation constants K₁ and K₂ for H₂S are 1×10^{-7} and 1×10^{-14} , respectively. What is the equilibrium constant for the reaction represented by the equation above at 25 °C?

- (A) $1 \times 10^{-15} / 5 \times 10^{-12}$
- (B) $5 \times 10^{-15} / 1 \times 10^{-7}$
- (C) $1 \times 10^{-7} / 5 \times 10^{-28}$
- (D) $5 \times 10^{-15} / 1 \times 10^{-28}$
- (E) $1 \times 10^{-28} / 5 \times 10^{-15}$

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