

Multiple Choice. Draw a circle around the letter that represents the best answer.

CALCULATORS CANNOT BE USED IN THE MULTIPLE CHOICE SECTION

1. A compound dissolves in water to form an aqueous solution that is a poor conductor of electricity. The compound may be composed of:
a) a strong acid b) a strong base c) a soluble salt d) covalent molecules
e) ions
2. Which of the following salts is insoluble in water?
a) Na_2S b) K_3PO_4 c) $\text{Pb}(\text{NO}_3)_2$ d) $\text{Cu}(\text{OH})_2$ e) All of these are soluble in water
3. Which of the following salts is (are) soluble in water?
I. lead(II) nitrate
II. lead(II) sulfide
III. lead(II) carbonate
a) I only b) II only c) III only d) I and II only e) I, II and III
4. Aqueous solutions of sodium sulfide and copper(II) chloride are mixed together. Which statement is correct?
a) Both NaCl and CuS precipitate from solution.
b) No precipitate forms.
c) CuS will precipitate from solution.
d) NaCl will precipitate from solution.
e) No reaction will occur.
5. The net ionic equation for the reaction of aqueous sodium hydroxide with aqueous nitric acid is:
a) $\text{Na}^+ + \text{OH}^- + \text{H}^+ + \text{NO}_3^- \rightarrow \text{Na}^+ + \text{NO}_3^- + \text{H}_2\text{O}$
b) $\text{Na}^+ + \text{OH}^- + \text{H}^+ + \text{NO}_3^- \rightarrow \text{NaOH} + \text{HNO}_3$
c) $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$
d) $\text{Na}^+ + \text{OH}^- \rightarrow \text{NaOH}$
e) None of these
6. For which one of the following acid solutions will 50.0 mL neutralize 50.0 mL of a 0.2 M $\text{Ba}(\text{OH})_2$ solution?
a) 0.1 M HCl b) 0.2 M HNO_3 c) 0.3 M HBr d) 0.1 M H_2SO_3 e) 0.2 M H_2SO_4
7. When $\text{NH}_3(\text{aq})$ is added to $\text{Cu}^{2+}(\text{aq})$, a precipitate initially forms. Its formula is:
a) $\text{Cu}(\text{NH}_3)_4^{2+}$ b) $\text{Cu}(\text{NO}_3)_2$ c) $\text{Cu}(\text{OH})_2$ d) $\text{Cu}(\text{NH}_3)_2^{2+}$ e) CuO
8. Which of the following are oxidation-reduction reactions?
I. $\text{PCl}_3 + \text{Cl}_2 \rightarrow \text{PCl}_5$
II. $\text{Cu} + 2 \text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2 \text{Ag}$
III. $\text{CO}_2 + 2 \text{LiOH} \rightarrow \text{Li}_2\text{CO}_3 + \text{H}_2\text{O}$
IV. $\text{FeCl}_2 + 2 \text{NaOH} \rightarrow \text{Fe}(\text{OH})_2 + 2 \text{NaCl}$
a) III b) IV c) I and II d) I, II, and III e) I, II, III, and IV

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**Peter J. Mikulecky,Michelle Rose
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