The Chemistry of Life ANSWER KEY

Section 2-1 The Nature of Matter (pages 35-39)

This section identifies the three particles that make up atoms. It also explains how atoms of the same element can have a different number of neutrons and describes the two main types of chemical bonds.

Atoms (page 35)

The basic unit of matter is called a(an) atom.
 It is the center of an atom, made up of protons and
 Describe the nucleus of an atom.
 neutrons.

3. Complete the table about subatomic particles. SUBATOMIC PARTICLES

Particle Charge Location in Atom Proton - Positive - Nucleus Neutron - Neutral - Nucleus Electron - Negative-Surrounding nucleus

4. Why are atoms neutral despite having charged particles?
Atoms have equal numbers of electrons and protons, and these subatomic particles have equal, but opposite, charges.

Elements and Isotopes (page 36)

5. What is a chemical element?

A chemical element is a pure substance that consists entirely of one type of atom.

6. What does an element's atomic number represent?

It represents the number of protons in an atom of the element.

Atoms of the same element that differ in the number of neutrons they contain are known as isotopes.

8. How are isotopes identified? Isotopes are identified by their mass number.

Why do all isotopes of an element have the same chemical properties? They have the same chemical properties because they have the same number of electrons.

Chemical Compounds (page 37)

10. What is a chemical compound?

A chemical compound is a substance formed by the chemical combination of two or more elements in definite proportions.

11. What does the formula for table salt indicate about that compound? The formula for table salt, NaCl, indicates that the elements from which table salt forms—sodium and chlorine—combine in a 1:1 ratio.

Chemical Bonds (pages 38-39)

Answers To The Chemistry Of Life

R Pring

Answers To The Chemistry Of Life:

The book delves into Answers To The Chemistry Of Life. Answers To The Chemistry Of Life is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Answers To The Chemistry Of Life, encompassing both the fundamentals and more intricate discussions.

- 1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Answers To The Chemistry Of Life
 - Chapter 2: Essential Elements of Answers To The Chemistry Of Life
 - Chapter 3: Answers To The Chemistry Of Life in Everyday Life
 - o Chapter 4: Answers To The Chemistry Of Life in Specific Contexts
 - ∘ Chapter 5: Conclusion
- 2. In chapter 1, this book will provide an overview of Answers To The Chemistry Of Life. The first chapter will explore what Answers To The Chemistry Of Life is, why Answers To The Chemistry Of Life is vital, and how to effectively learn about Answers To The Chemistry Of Life.
- 3. In chapter 2, the author will delve into the foundational concepts of Answers To The Chemistry Of Life. The second chapter will elucidate the essential principles that need to be understood to grasp Answers To The Chemistry Of Life in its entirety.
- 4. In chapter 3, this book will examine the practical applications of Answers To The Chemistry Of Life in daily life. The third chapter will showcase real-world examples of how Answers To The Chemistry Of Life can be effectively utilized in everyday scenarios.
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- 6. In chapter 5, this book will draw a conclusion about Answers To The Chemistry Of Life. The final chapter will summarize the key points that have been discussed throughout the book.
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