

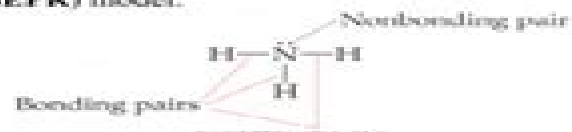
**STUDY GUIDE** *AP Chemistry*  
**CHAPTER NINE- MOLECULAR GEOMETRY AND BONDING THEORIES**  
**SECTIONS 9.1 THROUGH 9.6 ONLY**

### 9.1 Molecular Shapes

- Lewis structures give atomic connectivity: they tell us which atoms are physically connected to which atoms.
- The shape of a molecule is determined by its **bond angles**.
  - The angles made by the lines joining the nuclei of the atoms in a molecule are the bond angles.
- Consider  $\text{CCl}_4$ :
  - Experimentally we find all  $\text{Cl}-\text{C}-\text{Cl}$  bond angles are  $109.5^\circ$ .
  - Therefore, the molecule cannot be planar.
  - All Cl atoms are located at the vertices of a tetrahedron with the C at its center.
- In order to predict molecular shape, we assume that the valence electrons repel each other.
  - Therefore, the molecule adopts the three-dimensional geometry that minimizes this repulsion.
  - We call this model the **Valence Shell Electron Pair Repulsion (VSEPR)** model.

### 9.2 The VSEPR Model

- A covalent bond forms between two atoms when a pair of electrons occupies the space between the atoms.
  - This is a **bonding pair** of electrons.
  - Such a region is an **electron domain**.
- A **nonbonding pair** or *lone pair* of electrons defines an electron domain located principally on one atom.
- Example:  $\text{NH}_3$  has three bonding pairs and one lone pair.
- VSEPR predicts that the best arrangement of electron domains is the one that minimizes the repulsions among them.
  - The arrangement of electron domains about the central atom of a molecule is its **electron-domain geometry**.
  - There are five different electron-domain geometries:
    - linear (two electron domains), trigonal planar (three domains), tetrahedral (four domains), trigonal bipyramidal (five domains) and octahedral (six domains).
- The **molecular geometry** is the arrangement of the atoms in space.
  - To determine the shape of a molecule we must distinguish between lone pairs and bonding pairs.
    - We use the electron-domain geometry to help us predict the molecular geometry.
    - Draw the Lewis structure.
    - Count the total number of electron domains around the central atom.
    - Arrange the electron domains in one of the five geometries to minimize electron-electron repulsion.
    - Next, determine the 3-D structure of the molecule.
      - We ignore lone pairs in the molecular geometry.
      - Describe the molecular geometry in terms of the the bonded atoms.
    - Multiple bonds are counted as one electron domain.



Number of Electron Domains	Arrangement of Electron Domains	Electron-Domain Geometry	Preferred Bond Angles
2		Linear	$180^\circ$
3		Trigonal planar	$120^\circ$
4		Tetrahedral	$109.5^\circ$
5		Trigonal bipyramidal	$120^\circ$ $90^\circ$
6		Octahedral	$90^\circ$

#### The Effect of Nonbonding Electrons and Multiple Bonds on Bond Angles

We refine VSEPR to predict and explain slight distortions from "ideal" geometries.

- Consider three molecules with tetrahedral electron domain geometries:  $\text{CH}_4$ ,  $\text{NH}_3$ , and  $\text{H}_2\text{O}$ .
  - By experiment, the  $\text{H}-\text{X}-\text{H}$  bond angle decreases from C ( $109.5^\circ$  in  $\text{CH}_4$ ) to N ( $107^\circ$  in  $\text{NH}_3$ ) to O ( $104.5^\circ$  in  $\text{H}_2\text{O}$ ).
    - A bonding pair of electrons is attracted by two nuclei. They do not repel as much as lone pairs which are primarily attracted by only one nucleus.
  - Electron domains for nonbonding electron pairs thus exert greater repulsive forces on adjacent electron domains.
    - They tend to compress the bond angles.
    - The bond angle decreases as the number of nonbonding pairs increases.
  - Similarly, electrons in multiple bonds repel more than electrons in single bonds (e.g., in  $\text{Cl}_2\text{CO}$  the  $\text{O}-\text{C}-\text{Cl}$  angle is  $124.3^\circ$ , and the  $\text{Cl}-\text{C}-\text{Cl}$  bond angle is  $111.4^\circ$ ).

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## **Table of Contents Ap Chemistry Chapter 9 Study Guide**

1. Understanding the eBook Ap Chemistry Chapter 9 Study Guide
  - The Rise of Digital Reading Ap Chemistry Chapter 9 Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Ap Chemistry Chapter 9 Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ap Chemistry Chapter 9 Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Chemistry Chapter 9 Study Guide
  - Personalized Recommendations
  - Ap Chemistry Chapter 9 Study Guide User Reviews and Ratings
  - Ap Chemistry Chapter 9 Study Guide and Bestseller Lists
5. Accessing Ap Chemistry Chapter 9 Study Guide Free and Paid eBooks
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  - Adjustable Fonts and Text Sizes of Ap Chemistry Chapter 9 Study Guide
  - Highlighting and Note-Taking Ap Chemistry Chapter 9 Study Guide
  - Interactive Elements Ap Chemistry Chapter 9 Study Guide
- 8. Staying Engaged with Ap Chemistry Chapter 9 Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ap Chemistry Chapter 9 Study Guide
- 9. Balancing eBooks and Physical Books Ap Chemistry Chapter 9 Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ap Chemistry Chapter 9 Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ap Chemistry Chapter 9 Study Guide
  - Setting Reading Goals Ap Chemistry Chapter 9 Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ap Chemistry Chapter 9 Study Guide
  - Fact-Checking eBook Content of Ap Chemistry Chapter 9 Study Guide
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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