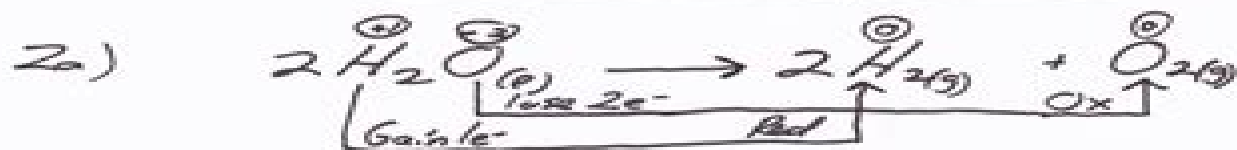
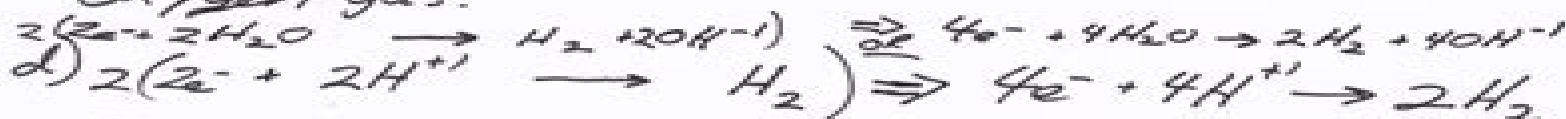


2005 AP Chemistry Test (Form B) Problem #2



b)
$$\frac{5.61 \text{ min}}{1} \times \frac{60 \text{ s}}{1 \text{ min}} \times \frac{.513 \text{ C}}{5} = 173 \text{ C}$$

- c) The balanced reaction shows that twice the amount hydrogen is produced per mole of oxygen produced. Therefore, twice the moles of hydrogen will produce twice the volume when compared to oxygen gas.



$$\frac{5.61 \text{ min}}{1} \times \frac{60 \text{ s}}{1 \text{ min}} \times \frac{.513 \text{ C}}{5} \times \frac{1 \text{ mol } e^-}{96,500 \text{ C}} \times \frac{2 \text{ mol } \text{H}_2}{4 \text{ mol } e^-} = 8.95 \times 10^{-4} \text{ mol } \text{H}_2$$

e) $n = 8.95 \times 10^{-4}$

$P = 1.0 \text{ atm}$

$T = 298 \text{ K}$

$R = .0821 \frac{\text{atm} \cdot \text{L}}{\text{mol} \cdot \text{K}}$

$V = ?$

$PV = nRT \Rightarrow V = \frac{nRT}{P}$

$$V = \frac{8.95 \times 10^{-4} \times .0821 \times 298}{1.0}$$

$V = .022 \text{ L} \quad (.0219 \text{ L})$

- f) The measured volume of hydrogen is greater than the dry volume measure in (e) because the gas was collected over water. The partial pressure of water in addition to the partial pressure creates a greater volume than what was calculated with only the pressure of hydrogen.

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Paul Foglino, Princeton Review (Firm)



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