Name	period	AP chemistry	Unit 5 worksheet
1.000			

- 1. Balance the following equations
  - a. 200+ O<sub>2</sub> → 200<sub>2</sub>
  - b. Cr(OH), + 3 HClO, → Cr(ClO<sub>2</sub>), + 3H<sub>2</sub>O
- 2. Write a balance equation to correspond to each description
  - When solid potassium chlorate is heated, it decomposes to form solid potassium chloride and oxygen gas.

2KCIO<sub>3</sub> → 2KCI + 3O<sub>2</sub>

 When liquid phosphorus trichloride is added to water it reacts to form aqueous phosphorous acid and aqueous hydrochloric acid

PCI, +3 H,O -> H,PO, + 3HCI

- 3. Write a balanced reaction for each of the following.
  - a. Uthium metal is burned in air

4LI + O. -> 2LI.O

Aluminum metal is added to a solution of copper (II) chloride

2AI + 3CuCl<sub>2</sub> → 2AICl<sub>3</sub> + 3Cu 2AI + 3Cu<sup>2+</sup> → 2AI<sup>2+</sup> + 3Cu

c. Manganese (II) nitrate solution is mixed with sodium hydroxide solution

 $Mn(NO_1)_1 + 2NaOH \rightarrow 2NaNO_1 + Mn(OH)_2$  $Mn^{2^n} + 2OH \rightarrow Mn(OH)_2$ 

d. The hydrocarbon hexane, CoHoo, is burned in excess oxygen

2C<sub>6</sub>H<sub>14</sub> + 19O<sub>2</sub> → 12 CO<sub>2</sub> + 14H<sub>2</sub>O

e. A sample of calcium carbonate is heated

CaCO, -> CaO + CO,

f. Silver nitrate solution is added to potassium chromate solution

 $2AgNO_1 + K_1CrO_1 \rightarrow Ag_2CrO_1 + 2KNO_1$  $2Ag^2 + CrO_1^2 \rightarrow Ag_2CrO_1$ 

g. Hydrogen peroxide decomposes

H,O. → 2H,O + O.

Solid calcium carbonate is reacted with sulfuric acid

 $CaCO_3 + H_2SO_4 \rightarrow CaSO_4 + H_2O + CO_2$  $CaCO_3 + 2H + \rightarrow Ca^{2*} + H_2O + CO_2$ 

i. Ammonium sulfate solution and potassium hydroxide solution are mixed together

 $[NH_u]_2SO_1 + 2KOH \rightarrow K_2SO_2 + 2NH_3 + 2H_2O$  $2NH_u^+ + 2OH \rightarrow 2NH_3 + 2H_2O$ 

j. Strontium hydroxide is mixed with hydrofluoric acid

 $Sr(OH)_2 + 2HF \rightarrow 2H_2O + SrF_2$  $2OH + 2HF \rightarrow 2H_2O + 2F$ 

# **Ap Chem Solutions Worksheet Answers**

John Moore, Richard H. Langley

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