

## AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of  $\text{O}_2$ ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\begin{array}{c} 21.5 \text{ g } \text{O}_2 \quad | \quad 1 \text{ mol } \text{O}_2 \quad | \quad 2 \text{ SO}_2 \quad | \quad 3 \text{ S} \quad | \quad 32.06 \text{ g S} \\ \hline 32.00 \text{ g } \text{O}_2 \quad | \quad 3 \text{ O}_2 \quad | \quad \text{SO}_2 \quad | \quad 1 \text{ mol S} \end{array} = 43.1 \text{ g S}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of  $\text{O}_2(g)$  are needed to produce 4.6 moles of  $\text{Fe}(s)$ ?A) 2.3 mole  $\text{O}_2$ B) 3.5 mole  $\text{O}_2$ C) 4.6 mole  $\text{O}_2$ D) 6.9 mole  $\text{O}_2$ E) 13.8 mole  $\text{O}_2$ 

$$\begin{array}{c} 4.6 \text{ mol Fe}(s) \quad | \quad 3 \text{ CO} \quad | \quad 1 \text{ O}_2 \\ \hline 2 \text{ Fe} \quad | \quad 2 \text{ CO} \end{array} = 3.45 \text{ mol } \text{O}_2$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of  $\text{H}_2\text{SO}_4$ . How much phosphoric acid will be produced?

A) 76.7 g

B) 51.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\begin{array}{c} 153 \text{ g } \text{Ca}_3(\text{PO}_4)_2 \quad | \quad 1 \text{ mol } \text{Ca}_3(\text{PO}_4)_2 \quad | \quad 2 \text{ H}_3\text{PO}_4 \\ \hline 310 \text{ g } \text{Ca}_3(\text{PO}_4)_2 \quad | \quad 1 \text{ Ca}_3(\text{PO}_4)_2 \quad | \quad 1 \text{ mol H}_3\text{PO}_4 \\ \hline 76.8 \text{ g H}_2\text{SO}_4 \quad | \quad 1 \text{ mol H}_2\text{SO}_4 \quad | \quad 2 \text{ H}_3\text{PO}_4 \\ \hline 64 \text{ g H}_2\text{SO}_4 \quad | \quad 3 \text{ H}_2\text{SO}_4 \quad | \quad 1 \text{ mol H}_3\text{PO}_4 \end{array} =$$

B

- 4.
- $\text{SO}_2$
- reacts with
- $\text{H}_2\text{S}$
- as follows:

When 7.50 g of  $\text{H}_2\text{S}$  reacts with 12.75 g of  $\text{SO}_2$ , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.C) 0.0216 moles of  $\text{H}_2\text{S}$  remain.D) 1.13 g of  $\text{H}_2\text{S}$  remain.E)  $\text{SO}_2$  is the limiting reagent.

$$\begin{array}{c} 7.50 \text{ g H}_2\text{S} \quad | \quad 1 \text{ mol H}_2\text{S} \quad | \quad 3 \text{ S} \quad | \quad 32 \text{ g S} \\ \hline 34.08 \text{ g H}_2\text{S} \quad | \quad 2 \text{ H}_2\text{S} \quad | \quad 1 \text{ mol H}_2\text{S} \end{array} = 10.89$$

$$\begin{array}{c} 12.75 \text{ g SO}_2 \quad | \quad 1 \text{ mol SO}_2 \quad | \quad 3 \text{ S} \quad | \quad 32 \text{ g S} \\ \hline 64 \text{ g SO}_2 \quad | \quad 1 \text{ SO}_2 \quad | \quad 1 \text{ mol S} \end{array} = 19.13$$

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**The Princeton Review**



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