

POGIL: Periodic Table Trends

(Adapted from: <https://www.khanacademy.org/a/periodic-table-trends>)

Why?

Chemists have been able to predict trends in the behavior of atoms in atoms, molecules, and ionic compounds, and to predict properties of these substances. Chemists understand, for example, reacting and bonding behavior in chemical reactions and together the rules within each atom's bound the properties of the elements based on the atom. They predict the trends in properties of the elements and the trends in these properties in the periodic table.



Periodic Table

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1. What happens to the number of valence electrons as you go from a group to the periodic table? Does the number of valence electrons increase or decrease?

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Ap Chem Periodic Trends Pogil

Robert Thomas Sanderson



Ap Chem Periodic Trends Pogil:

Chemistry and Our Universe , Return to the periodic table introduced in Lecture 1 to practice predicting properties of elements based on their electronic structure Then witness what happens when three different alkali metals react with water Theory forecasts a pronounced difference in the result Is there *Chemical Periodicity* Robert Thomas Sanderson,2013-04-20

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The Periodic Table Eric R. Scerri,2019-07-25 The periodic table of elements first encountered by many of us at school provides an arrangement of the chemical elements ordered by their atomic number electron configuration and recurring chemical properties and divided into periodic trends In this Very Short Introduction Eric R Scerri looks at the trends in properties of elements that led to the construction of the table and shows how the deeper meaning of the table s structure gradually became apparent with the development of atomic theory and in particular quantum mechanics which underlies the behaviour of all of the elements and their compounds This new edition publishing in the International Year of the Periodic Table celebrates the completion of the seventh period of the table with the ratification and naming of elements 113 115 117 and 118 as nihonium moscovium tennessine and oganesson Eric R Scerri also incorporates new material on recent advances in our understanding of the origin of the elements as well as developments concerning group three of the periodic table ABOUT THE SERIES The Very Short Introductions series from Oxford University Press contains hundreds of titles in almost every subject area These pocket sized books are the perfect way to get ahead in a new subject quickly Our expert authors combine facts analysis perspective new ideas and enthusiasm to make interesting and challenging topics highly readable

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New Ideas in Chemistry from Fresh Energy for the Periodic Law Henry

Bent, 2006-09-05 New Ideas calls to mind Aristotle's synopsis of the Iliad and the Odyssey. Woman abducted. Long war. One guy has a hard time getting home. End of story. The rest is episodes. Similarly here. Chemical capture of the Left Step Periodic Table. One element finds a new home. The noblest of the noble gases is not a Noble Gas. End of story. The rest is novel consequences of the Noble Gas Conclusion. Among them overlooked Rules of Triads, Block Sizes, and Full Shells, overlooked block to block trends, and a correspondence between elements' ordinal numbers in their Groups and orbital's radial quantum numbers. Recognition that Pauli's explanation of Periodicity's magic numbers 2, 8, 18 got the right answer, the Pauli Exclusion Principle, for the wrong reason. New Ideas ends with suggestions for streamlining the teaching of the mole concept, chemical bonding, and thermodynamics in order to provide room in the chemistry curriculum for a more thorough treatment of Periodic System Systematics.

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